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Near-Ambient Pressure XPS and NEXAFS Study of a Superbasic Ionic Liquid with CO₂

Supporting Information

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XPS Spectra

Figure S1 shows the survey spectra of a thin film of [P₆₆₆₁₄][benzim] as presented, while exposed to 10⁻⁶ mbar CO₂, 1 mbar CO₂, 5 mbar CO₂, and after exposure. A survey scan of the clean TiO₂ (110) crystal was taken prior to deposition of the IL. These were all taken with a photon energy of 1200 eV. The assignments and binding energies (BEs) of fitted components across the C 1s, N 1s, O 1s and P 2p regions is summarised in Table S1.

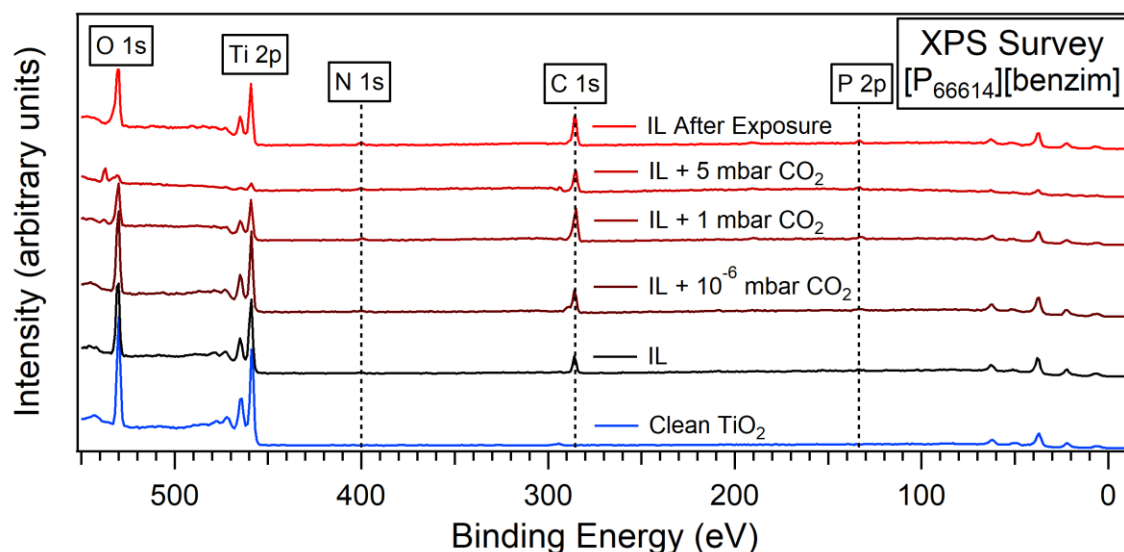


Figure S1. XPS survey spectra of [P₆₆₆₁₄][benzim] subject to various exposures of CO₂. Highlighted are the peaks associated with atoms in the TiO₂ (110) crystal substrate (O 1s, Ti 2p) and the IL (N 1s, C 1s, P 2p).

Narrow scans of the Ti 2p region for the IL [P₆₆₆₁₄][benzim] are shown in Figure S2 before exposure to CO₂; upon exposure to 10⁻⁶ mbar, 1 mbar and 5 mbar of CO₂; and after exposure. The region consists of two peaks at 458.7 and 464.5 eV which occur due to spin-orbit coupling, assigned as 2p_{3/2} and 2p_{1/2}, respectively. The spectra show very little change through the exposure stages, except for increasing noise with higher pressures of CO₂ due to increased attenuation of photoelectrons.

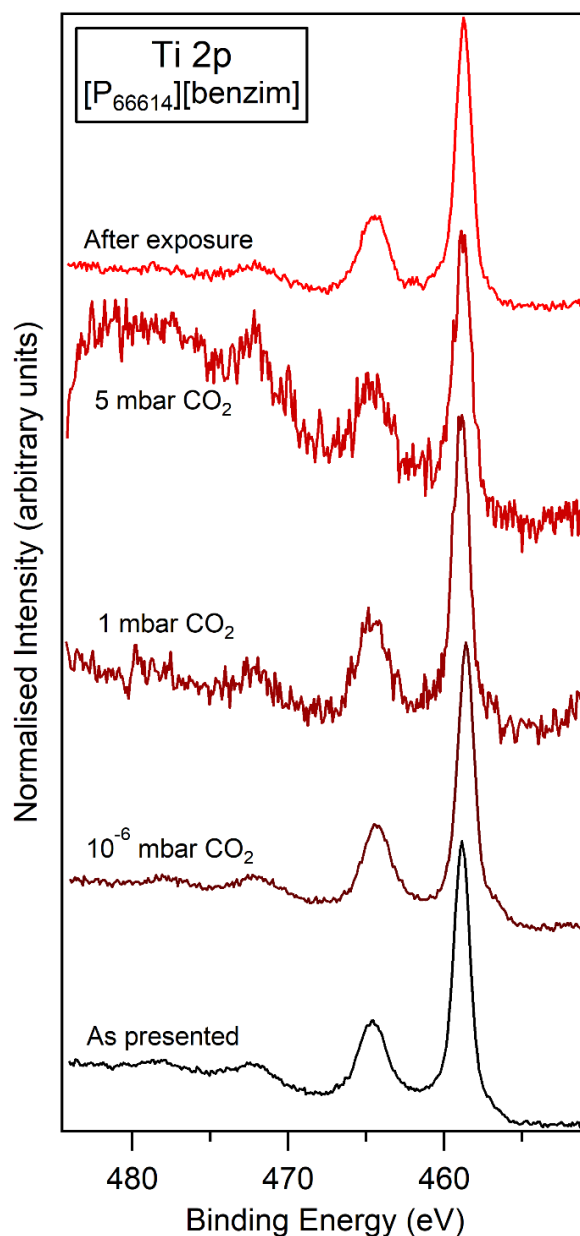


Figure S2. Ti 2p region for multilayer [P₆₆₆₁₄][benzim] as presented, exposed to 10⁻⁶ mbar CO₂, 1 mbar CO₂, 5 mbar CO₂, and after exposure (bottom to top, respectively).

O 1s spectra taken over the same exposure stages are shown in Figure S3. When the IL is exposed to higher pressures of CO₂ (1 mbar and 5 mbar), the spectra can be fitted with four components at binding energies of 530.0, 531.1, 532.2 and 533.3 eV. Note that the intense gas-phase CO₂ peak at 537.2 eV is not shown in Figure S3 in order to make the region clearer. The component at 530.0 eV is attributed to O atoms in the TiO₂ lattice and is detectable

because the IL film is sufficiently thin (approximately 5 nm).¹ We suspect there is a mixture of protonated and unprotonated carbamate species in the sample, therefore, the component at 531.1 eV is attributed to unprotonated carbamate, while the remaining components at 532.2 and 533.3 eV are attributed to C = O and C – OH in the protonated carbamate group, respectively.^{2–4} Protonation of the anion is likely to arise due to contamination of water vapour. The TiO₂ lattice O component at 530.0 eV reduces in intensity with increasing CO₂ pressure. Upon exposure to the highest pressure of 5 mbar CO₂, the C = O and C – OH components contribute to a pronounced protonated carbamate feature. After removal of CO₂, the protonated carbamate components are still present (top spectrum of Figure S3), however, in the IL as presented and IL + 10⁻⁶ mbar CO₂ stages the C – OH component is too small to be fitted.

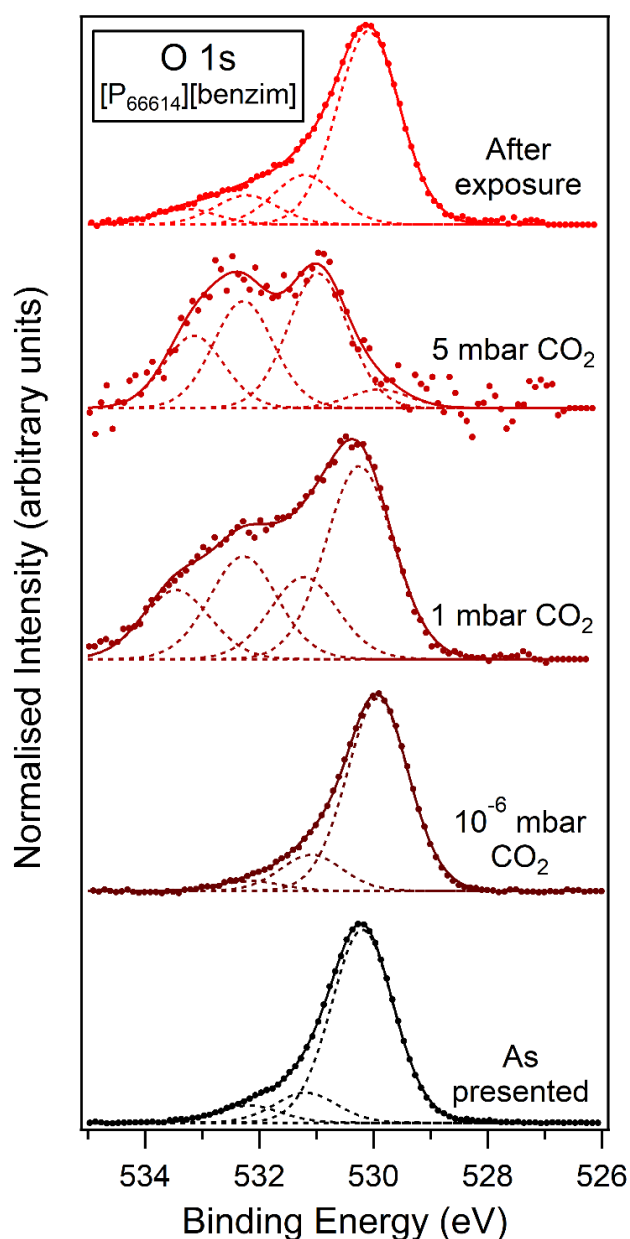


Figure S3. O 1s region recorded at a photon energy of 830 eV for multilayer [P₆₆₆₁₄][benzim] as presented, while exposed to various pressures of CO₂, and after exposure.

The O 1s spectra in Figure S3 have been normalised to the intensity of the TiO₂ lattice O peak at 530.0 eV. However, there is a dramatic relative decrease in the 530.0 eV component of the IL + 5 mbar CO₂ spectrum as the IL absorbs more CO₂. Therefore, this spectrum was normalised such that the intensities of high BE components at 532.2 and 533.3 eV have the same intensities as those for 1 mbar CO₂.

Region	Binding Energy (eV) \pm 0.1 eV	Assignment
C 1s	285.0	C _{aliphatic}
	285.9	C _{hetero}
	286.8 \rightarrow 287.1	C _{anion}
	289.0 ^a	Carbamate
	293.3	Gas-Phase CO ₂
N 1s	398.6	N ^{1,2}
	400.0	N ^{2,r}
	401.1	N ^{1,r}
O 1s	530.0 ^a	TiO ₂ lattice O
	531.1	Carbamate
	532.2	C = O ^b
	533.3 ^a	C – OH ^b
	537.2	Gas-Phase CO ₂
P 2p	130.4	THP 2p _{3/2}
	131.3	THP 2p _{1/2}
	132.4 \rightarrow 132.8	[P ₆₆₆₁₄] ⁺ 2p _{3/2}
	133.3 \rightarrow 133.7	[P ₆₆₆₁₄] ⁺ 2p _{1/2}
	135.6	Phosphine oxide 2p _{3/2}
	136.5	Phosphine oxide 2p _{1/2}

Table S1. Assignments and BEs of fitted components in the C 1s, N 1s, O 1s and P 2p regions for [P₆₆₆₁₄][benzim] before exposure, while exposed to CO₂ at various pressures, and after exposure. These were taken at photon energies of 585, 700, 830, and 430 eV, respectively, for each of the above regions. The arrow (\rightarrow) denotes a chemical shift due to the absorption of CO₂. ^a BEs differ by \pm 0.2 eV. ^b Within the protonated carbamate group.

Grazing Emission XPS Study

XPS spectra of [P₆₆₆₁₄][benzim] were taken at grazing emission (GE, 40° from normal emission) over the C 1s region as shown in Figure S4, providing more surface-sensitive measurements. There is little change in the overall shape of the C 1s peak when taken at GE compared to normal emission (NE) except for a slight reduction in the intensity of the C_{hetero} component at 285.9 eV. From this we can deduce that the concentration of C_{aliphatic}, C_{hetero} and C_{anion} species does not vary significantly throughout the sample depth. The peaks at 293 and 296 eV are assigned as K 2p_{3/2} and K 2p_{1/2}, respectively. They become more pronounced for GE, indicating

surface contamination with K. Shake-up satellites may also manifest at similar BEs but the presence of a feature at approximately 295 eV in the clean TiO₂ survey spectrum would suggest this is K contamination.

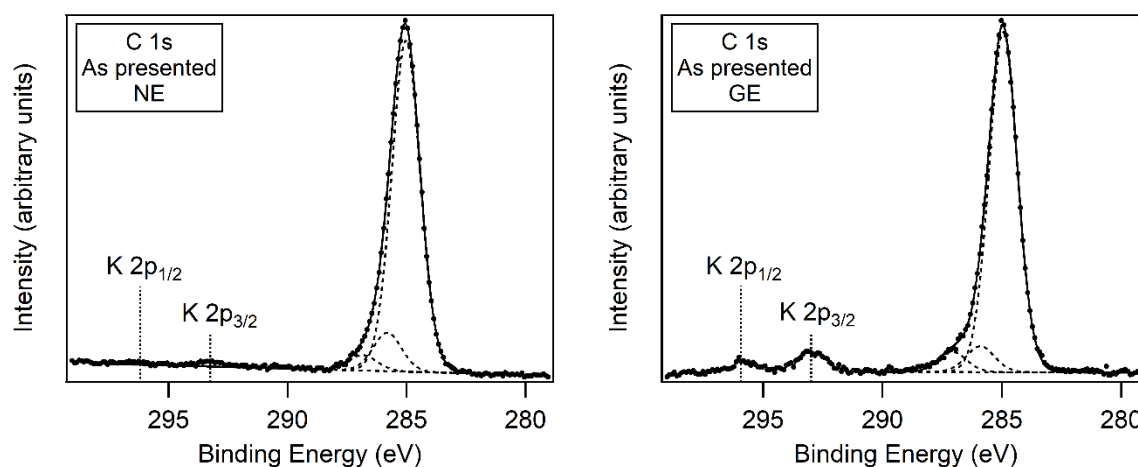


Figure S4. C 1s normal emission (NE) and grazing emission (GE) XPS spectra of [P₆₆₆₁₄][benzim].

GE measurements over the N 1s region were taken before exposure to CO₂ and compared to those taken at NE as shown in Figure S5. This provides an insight into how the concentration of the N^{1,r} (401.1 eV) and N^{2,r} (400.0 eV) reaction species changes with depth into the sample. Compared to NE, the N 1s peak taken at GE in Figure S5 has a more intense N^{1,2} (398.6 eV) component relative to the two reaction-based components. This suggests that there is a greater concentration of reacted anions in the deeper layers of the thin film.

As discussed by Henderson et al., the anion can also react with ambient water vapour, resulting in components at similar BEs of 397.5, 398.5 and 400.0 eV.⁵ Our results in Figure 1(c) show a significant increase in BE for these components for multilayer versus bulk deposition. Wagstaffe et al. reported a dependence of peak position on overlayer thickness, finding that components undergo a downward shift in BE when deposited as a thin film compared to bulk deposition.⁶ Coverage-dependent chemical shifts have been found for thin films of the IL [C₁C₁Im][Tf₂N] deposited on Au (111) via physical vapour deposition by Cremer et al.⁷ They reported a similar dependence to Wagstaffe et al. with the addition that the peak positions for thin films of increasing thicknesses did not tend to those of a bulk film but exceeded them by as much as 0.2 eV (4.5 monolayer film). This may explain the higher BEs reported here for our thin film in Figure 1(c) compared to previous reports for bulk thicknesses.

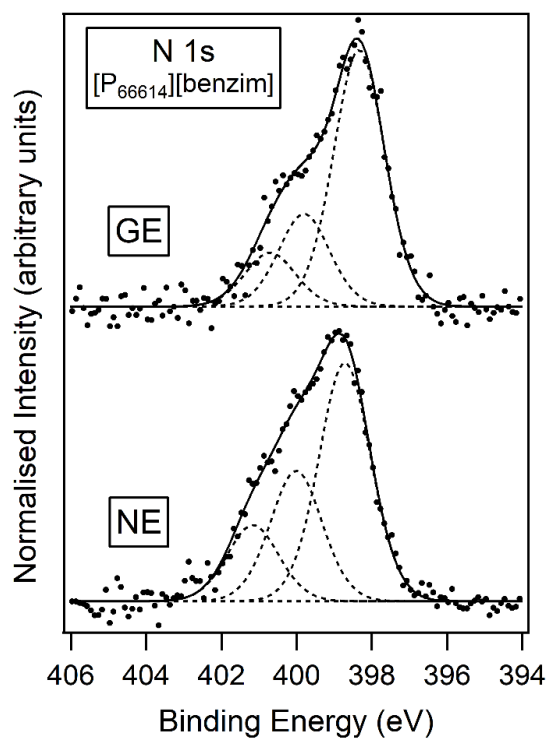


Figure S5. *N 1s spectra of multilayer $[P_{66614}][benzim]$ (as presented) recorded at a photon energy of 700 eV at normal and grazing emissions.*

GE measurements over the P 2p region (Figure S6) show a more intense 133 eV $[P_{66614}]^+$ peak at GE compared to NE before exposure to CO_2 , suggesting there is a greater concentration of $[P_{66614}]^+$ cations at the surface compared to the bulk.

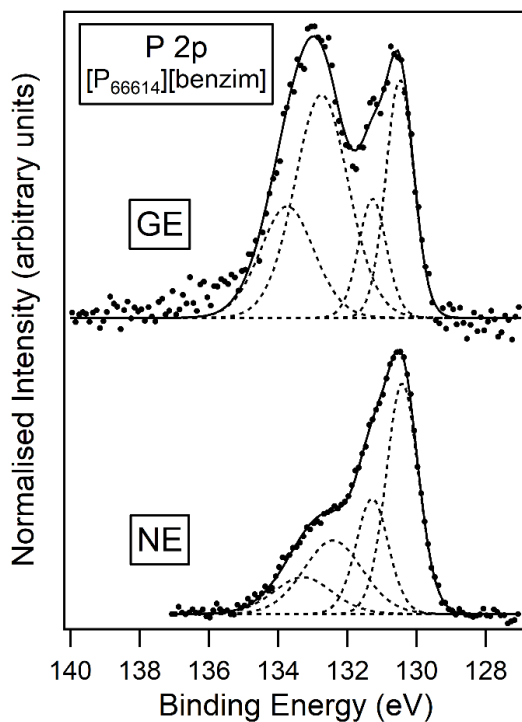


Figure S6. *P 2p spectra of multilayer $[P_{66614}][benzim]$ (as presented) recorded at a photon energy of 430 eV at normal and grazing emissions.*

GE measurements were also taken over the O 1s region and compared to NE measurements of the IL thin film and a clean TiO₂ (110) substrate, shown in Figure S7. There is no change in the spectra of the IL thin film when taken at GE compared to NE.

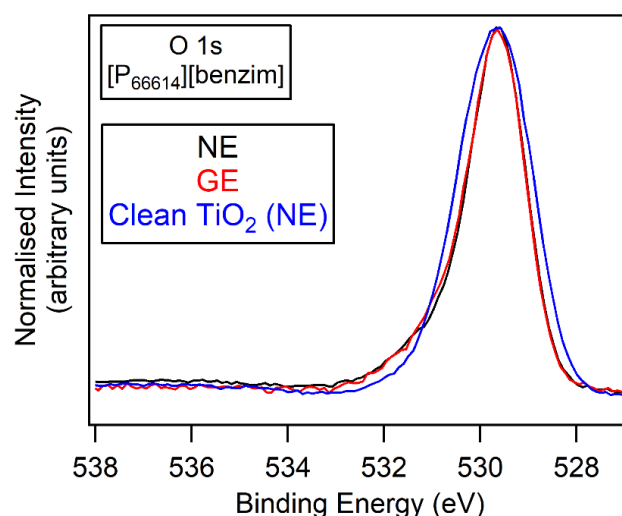


Figure S7. O 1s normal and grazing emission of [P₆₆₆₁₄][benzim]. This is compared to clean TiO₂ (110) at normal emission.

NEXAFS CO₂ Gas Correction

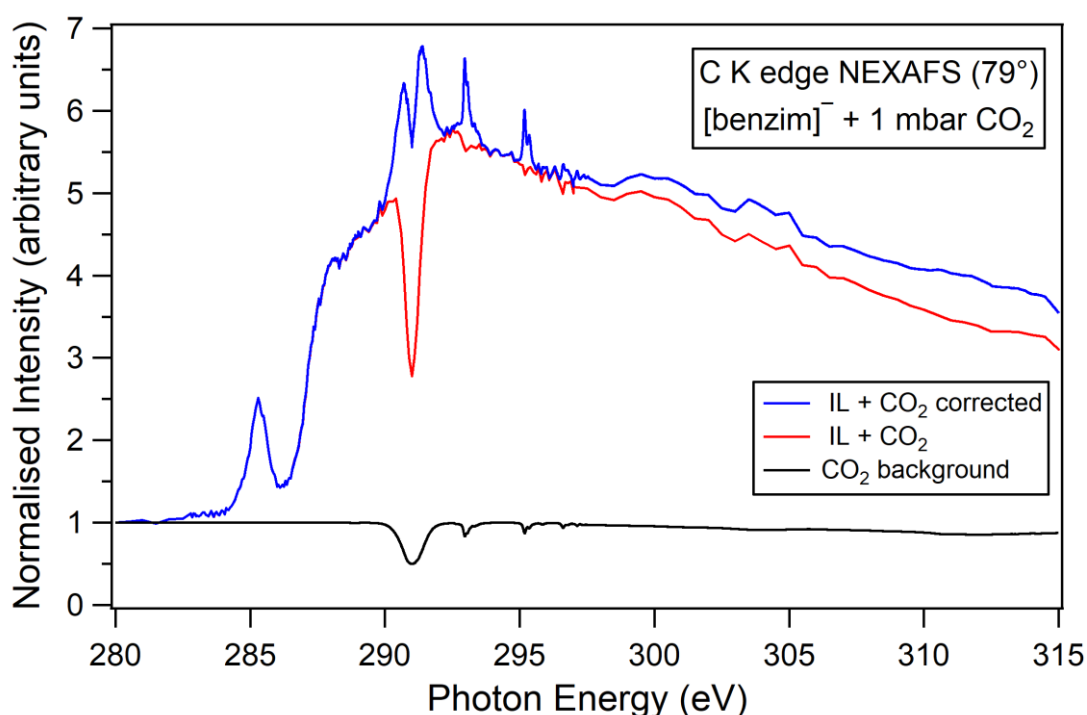


Figure S8. C K edge NEXAFS of [P₆₆₆₁₄][benzim] exposed to 1 mbar CO₂. The experimental data (red) was divided by the background gas-phase CO₂ signal (black), resulting in a CO₂-corrected spectrum (blue).

The CO₂ gas in the NAP cell absorbs incoming X-rays at certain energies, introducing dips in the C K edge NEXAFS spectrum for [P₆₆₆₁₄][benzim] exposed to 1 mbar CO₂ (red line) in Figure S8. This was divided by the background signal due to gas-phase CO₂ (black line), which

removes the large dip at 291 eV but has the adverse effect of introducing small spikes into the background-corrected spectrum (blue line).

In NEXAFS measurements of the IL, we also corrected for the TiO_2 substrate onto which the IL is deposited. This was done by dividing the scan of the IL by the scan of the clean TiO_2 substrate. The example given in Figure S9 shows N K edge measurements of the IL exposed to 1 mbar of CO_2 (gold line), the clean TiO_2 signal (black), and the background-corrected IL + CO_2 signal obtained through division of the TiO_2 spectrum.

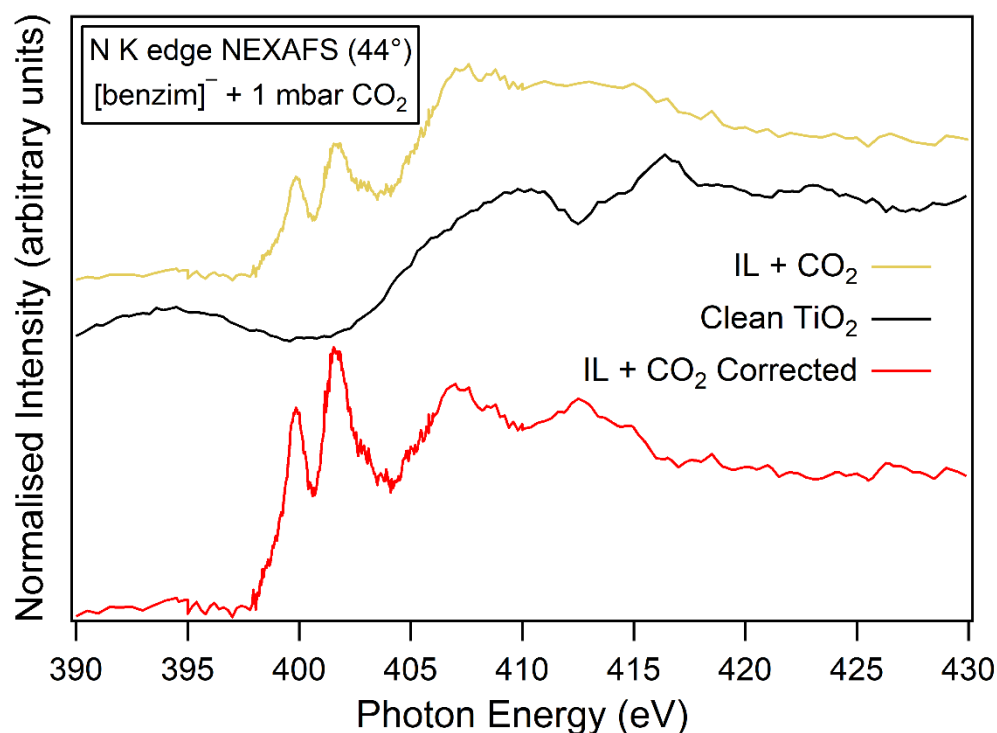


Figure S9: N K edge NEXAFS of $[\text{P}_{66614}][\text{benzim}]^-$ exposed to 1 mbar CO_2 with a removed TiO_2 background signal. The experimental data (gold) was divided by the clean TiO_2 substrate signal (black), resulting in a background-corrected IL + CO_2 spectrum (red).

Experimental N K edge NEXAFS

NEXAFS can be used to determine the orientation of $[\text{benzim}]^-$ anions in the IL thin film by comparing changes in the relative intensities of the π^* and σ^* peaks at different X-ray angles of incidence. π^* and σ^* resonances are maximised when the electric field vector of the polarised synchrotron light is parallel to the spatial orientation of the final state molecular orbital, i.e. resonances are maximum when the incident X-rays are perpendicular to the molecular orbital.⁸

N K edge NEXAFS of the IL as presented in Figure S10 has a feature at approximately 400 eV consisting of two peaks at photon energies of 399.9 and 401.7 eV. These peaks are attributed to $\text{N } 1s \rightarrow \pi^*$ transitions. There is a broader peak at 408.0 eV attributed to $\text{N } 1s \rightarrow \sigma^*$ transitions. There is no definitive evidence of realignment of $[\text{benzim}]^-$ anions in the as presented film in Figure S10 due to the noisy nature of the spectra. Walsh et al. reported a split π^* peak in multilayer benzimidazole on Cu (100) at similar photon energies, and also for

the similarly-structured benzotriazole on Cu (100).⁹ Grillo et al. reported a split π^* peak at different positions for benzotriazole on Au (111), revealing that the positions of these split π^* resonances vary depending on the substrate onto which the molecules are deposited.¹⁰ Splitting of the π^* peak is due to initial state effects, arising from chemical shifts between different environments of N atoms in the imidazolidine ring.^{1,11} The ions in these examples above are not charged and the N atoms would, therefore, be inequivalent, resulting in a split π^* peak. However, our DFT simulations of the isolated [benzim]⁻ anion (see Figure S14) show that both N atoms contribute to the split π^* peak equally, implying they share the same chemical environment. The origin of the split π^* peak for [benzim]⁻ will be discussed below.

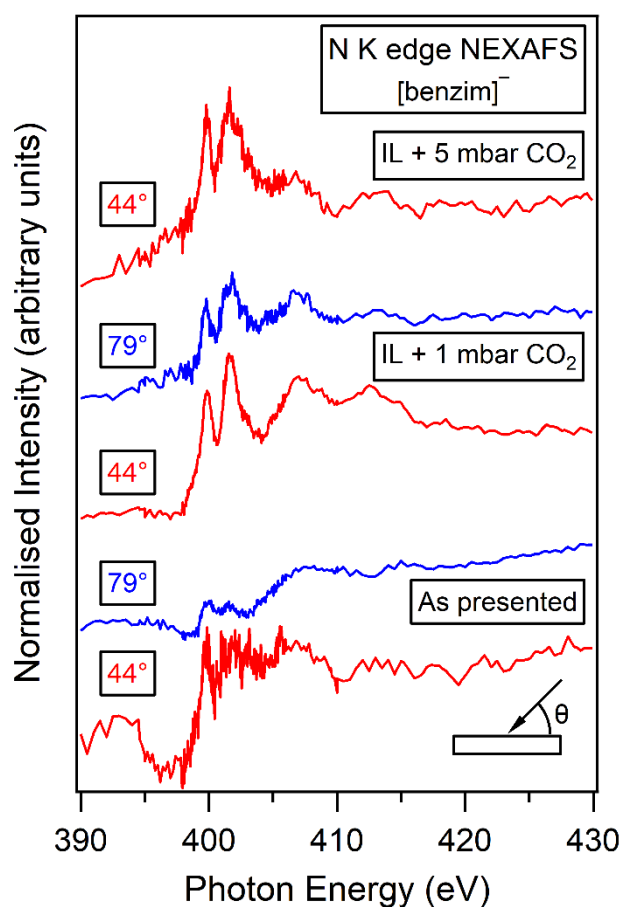


Figure S10. N K edge NEXAFS spectra of [P₆₆₆₁₄][benzim] as presented (bottom), when exposed to 1 mbar of CO₂ (middle) and when exposed to 5 mbar of CO₂ (top), for the X-ray beam incident at angles of 44° and 79° from the IL surface.

To investigate the realignment of [benzim]⁻ anions upon absorption of CO₂ the IL was subject to two pressures of CO₂, 1 mbar and 5 mbar, as shown in Figure S10. Both the π^* and σ^* features become more prominent than those for the IL as presented, particularly the split π^* peak. However, there is no obvious change in the relative intensity of these π^* and σ^* peaks due to the noisy nature of the data. As found with 1 mbar CO₂, when the IL is exposed to the higher 5 mbar CO₂ pressure we still see a split π^* peak. However, the higher-energy π^* peak at 401.7 eV is less intense relative to the lower-energy peak at 399.9 eV compared to 1 mbar CO₂. The lower-energy peak remains unchanged with CO₂ pressure, while all other features reduce in relative intensity with increased CO₂ pressure. At a 5 mbar pressure the π^* peaks

are much more intense relative to the σ^* peak, suggesting that the ring of the anion orients more in-plane than it does when exposed to a lower pressure of 1 mbar (i.e. the anion ring lies closer to the surface with higher CO₂ pressures). We do not have measurements at another angle to verify this, however, from results in Figure S15 that will be discussed later, we find that N¹ provides the greatest contribution to the σ^* peak. This is the site at which carbamate forms, and we would therefore expect a higher intensity σ^* peak upon exposure to higher pressures of CO₂ (5 mbar).

NEXAFS DFT Simulations

Density Functional Theory (DFT) simulations were carried out in the software StoBe-deMon, prior to which the isolated anion of the IL was geometry optimised in Avogadro. These simulations are not angle-resolved, meaning they are representative of anions being randomly oriented with respect to the incident X-ray beam. The final DFT NEXAFS spectrum is the sum of the individual simulated spectra for each target centre in the anion (e.g. each C atom for C K edge simulations).

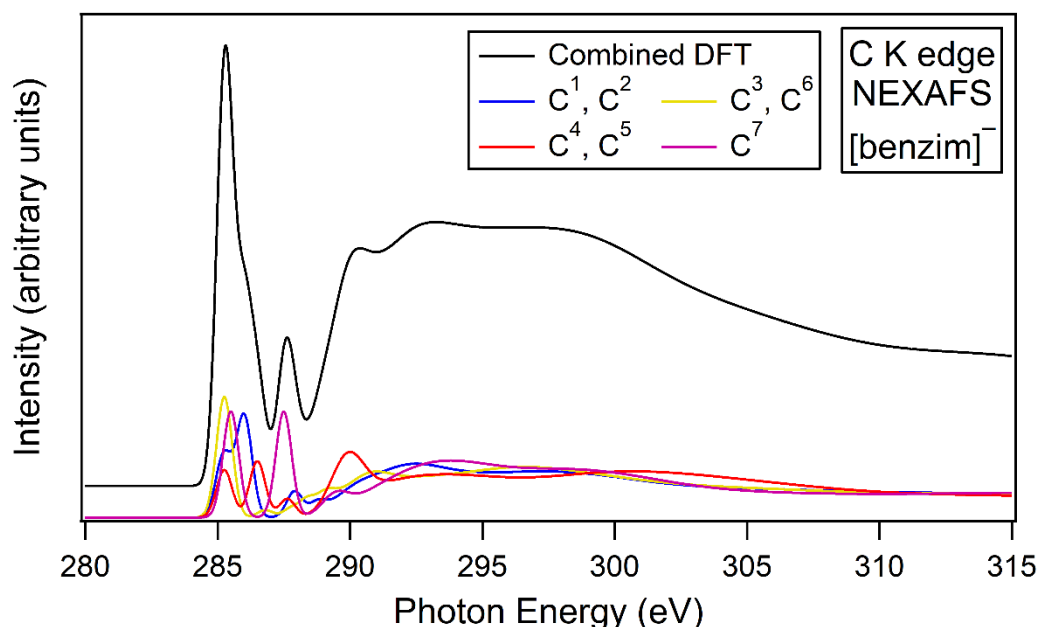


Figure S11. Simulated C K edge NEXAFS spectrum for [P₆₆₆₁₄][benzim]. The final DFT spectrum (black) is the sum of the seven spectra from atoms C¹ – C⁷.

The simulated C K edge NEXAFS for the IL isolated [benzim][−] anion is shown in Figure S11. The breakdown of the combined DFT spectrum (black line) indicates that there are four distinct environments of C atoms in the [benzim][−] anion: C^{1,2}; C^{3,6}; C^{4,5}; C⁷ (see Figure 1(a) for assignments). Comparing this to the experimental spectra in Figure 3, the π^* peak at 285.3 eV in the experimental spectrum is largely attributed to transitions from atoms C³, C⁶ and C⁷, with a smaller contribution from C¹, C², C⁴ and C⁵. The broadening of the 285.3 eV π^* peak in the DFT spectrum is due to transitions from atoms C¹, C², C⁴ and C⁵ in the [benzim][−] anion. The π^* peak at 287.9 eV in the experimental spectrum is largely attributed to transitions from atom C⁷. This peak occurs at higher photon energies because it is located between the two electronegative N atoms, consistent with reports for imidazolium-based ILs by Ehlert et al.¹¹

Although in their case the splitting of the π^* peak is much narrower, with a split of 0.3 eV, compared to a wider separation of 2.6 eV here for $[\text{benzim}]^-$. The π^* peak at 290.2 eV is largely due to transitions from atoms C^4 and C^5 , which also have neighbouring N atoms.

Figure S12 shows $[\text{benzim}]^- + \text{CO}_2$ over the C K edge. The combined DFT spectrum is the sum of all C atoms in the $[\text{benzim}]^-$ anion (C^{1-7}) and the C atom in the carbamate species (C^8). When reacted with CO_2 , the four environments of C atoms in $[\text{benzim}]^-$ expand into eight: C^{1-7} for each C atom in the $[\text{benzim}]^-$ anion, and C^8 in the carbamate species. The pre-edge peaks at approximately 282 and 284 eV are attributed to atom C^4 near the carbamate. We believe there may be protonated carbamate groups present in the sample so we have also calculated NEXAFS spectra for IL + COOH. The pre-edge features are not present in the protonated carbamate calculation (Figure S13), so they are most likely a result of the interaction between C^4 and the COO^- moiety. The lower-energy π^* peak at 285.3 eV in the experimental spectrum is mostly attributed to transitions from atoms C^1 , C^2 , C^3 and C^6 , those furthest from the azole side of the $[\text{benzim}]^-$ anion. Both the broadening of this peak towards higher photon energies and the π^* peak at 287.2 eV can be largely attributed to the remaining atoms nearer the carbamate: C^4 , C^5 and C^7 (with smaller contributions from C^1 and C^2). We also see that C^8 in the carbamate moiety does not contribute to any of the three π^* peaks or the σ^* peak at 291.1 eV, but rather contributes solely to σ^* resonances greater than 295 eV. This may be because the negative charge was assigned to the singly bonded O atom in the carbamate group (in the software used the negative charge has to be assigned to one atom rather than delocalised).

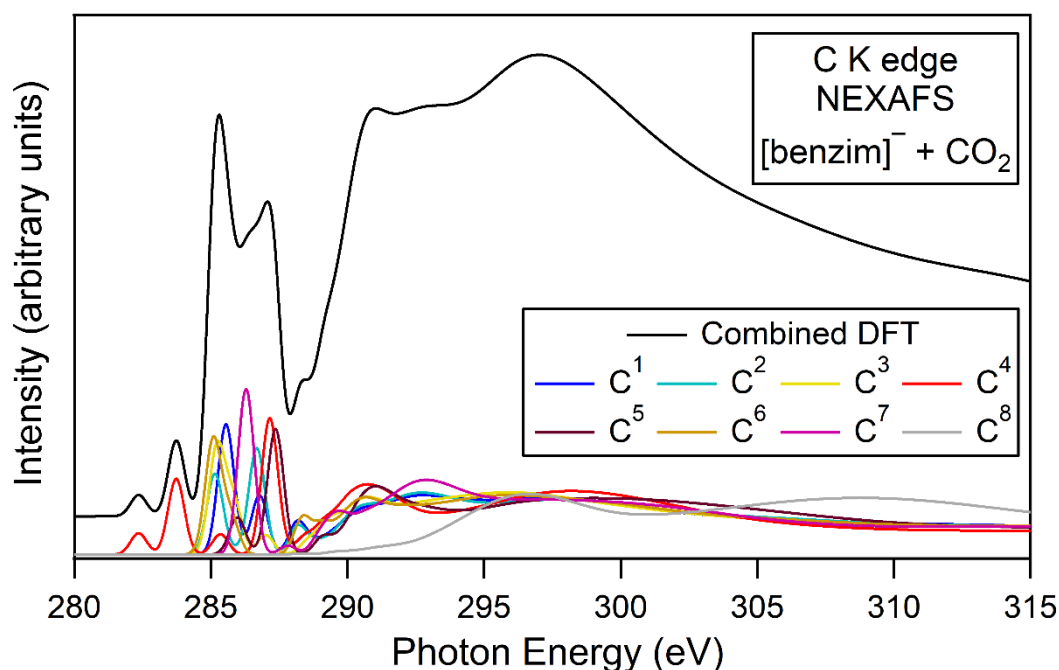


Figure S12. Simulated C K edge NEXAFS spectrum for $[\text{P}_{66614}][\text{benzim}] + \text{CO}_2$. The final DFT spectrum (black) is the sum of the eight spectra from atoms $\text{C}^1 - \text{C}^8$.

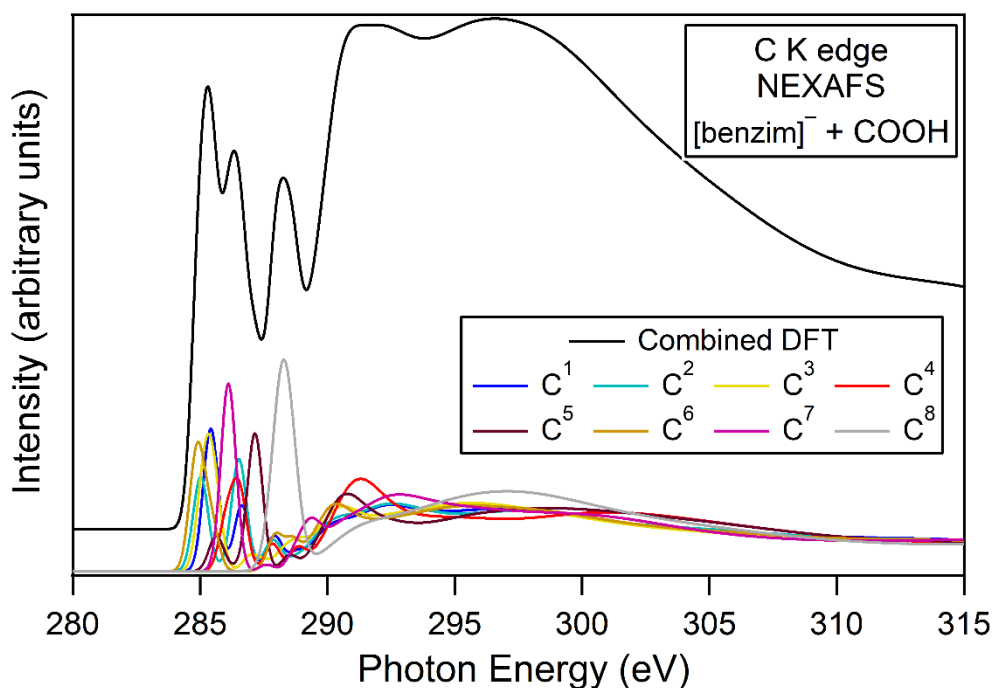


Figure S13: Simulated C K edge NEXAFS spectrum for $[P_{66614}][benzim] + COOH$. The final DFT spectrum (black) is the sum of the eight spectra from atoms $C^1 - C^8$.

DFT simulations were also carried out for $[benzim]^-$ over the N K edge as shown in Figure S14. In this case, atoms N^1 and N^2 have the same spectra, contributing to the final DFT spectrum equally. This is because an isolated $[benzim]^-$ anion is simulated (in which N^1 and N^2 are chemically equivalent) and the cation's influence on the anion is not considered. Mercy et al. have shown that the interaction between a $[P_{3333}]^+$ cation and a $[124Triz]^-$ anion results in preferential reactivity at N^1 , before which the reaction enthalpy was equivalent at both N sites.¹² Although it is possible that the influence of the $[P_{66614}]^+$ cation may affect the reaction enthalpy and/or chemical bonding of the N atoms, it is not expected to be significant.

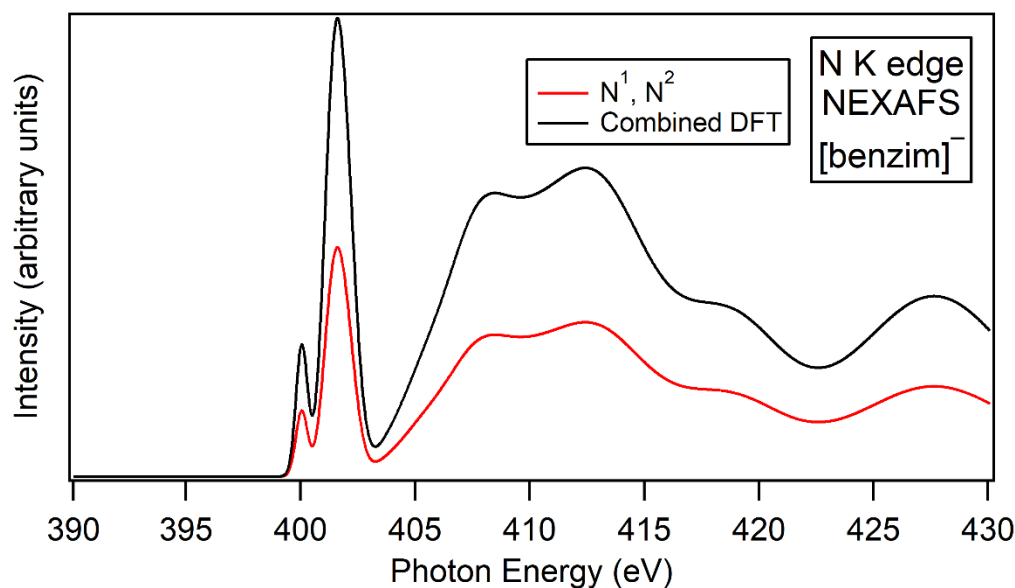


Figure S14. Simulated N K edge NEXAFS spectrum for $[P_{66614}][benzim]$. The final DFT spectrum (black) is the sum of the two spectra from atoms N^1 and N^2 (red).

When the IL absorbs CO_2 (Figure S15) we see that the spectra for N^1 and N^2 are no longer equivalent due to the formation of carbamate at $N^{1,r}$. The formation of this carbamate species destroys the $N^1 \rightarrow \pi^*$ resonances. Therefore, in the final DFT spectrum, we see that the π^* peak is solely attributed to transitions from $N^{2,r}$ (N^2 in the CO_2 -reacted anion). The spectra for N^2 before exposure (Figure S14) and upon exposure to CO_2 (Figure S15, $N^{2,r}$) are also different, showing that the formation of carbamate at $N^{1,r}$ influences π^* transitions from $N^{2,r}$.

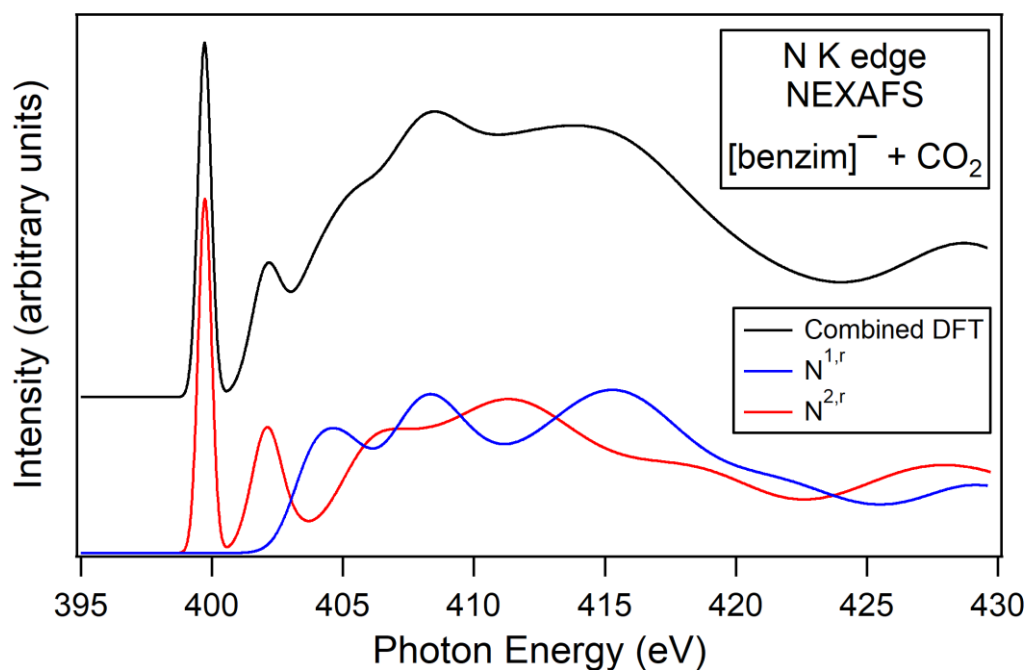


Figure S15. Simulated N K edge NEXAFS spectrum for $[P_{66614}][benzim] + CO_2$. The final DFT spectrum (black) is the sum of the two spectra from atoms $N^{1,r}$ (red) and $N^{2,r}$ (blue).

Theoretical DFT spectra were compared to experimental spectra over the N K edge for the IL as presented and when exposed to CO₂, as shown in Figure S16. For the IL as presented, the lower- and higher-energy π^* peaks appear at photon energies of 399.9 and 401.4 eV, respectively, similar to those in the experimental spectra in Figure S10 (399.9 and 401.7 eV). The σ^* peak of the simulated NEXAFS spectrum occurs at the same energy as that of the experimental spectrum (408.0 eV). Thomason et al. studied aqueous imidazole using experimental and theoretical NEXAFS over the N K edge and found a split π^* feature with two peaks at 400.2 and 401.9 eV.¹³ This is similar to the π^* peaks of the experimental and simulated NEXAFS of our benzimidazole-based IL in Figures S10 and S16, respectively. Thomason et al. suggest that the lower- and higher-energy π^* peaks are assigned to transitions from atoms N² and N¹ in the imidazole, respectively. However, in our case, N¹ and N² are found to contribute in equal parts to the simulated NEXAFS spectrum, as shown in Figure S14. The difference in results may be due to a number of reasons. The main cause of disparity may be due to differences in the chemical environment of the N atoms (N atoms are chemically equivalent in benzimidazole but not in imidazole) and charge of the two systems.

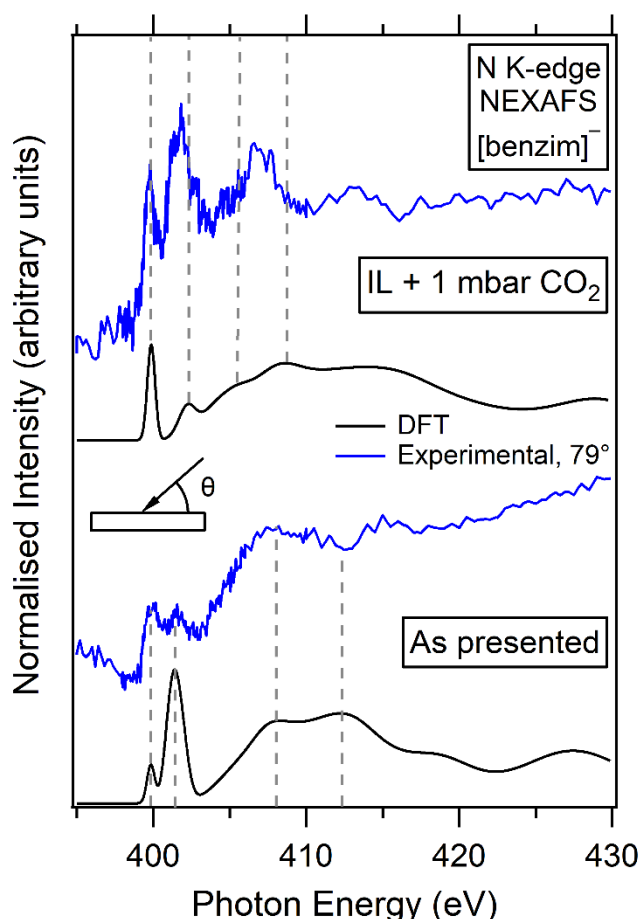


Figure S16. Experimental and simulated (DFT) N K edge NEXAFS of $[P_{66614}][\text{benzim}]^-$ as presented and when exposed to 1 mbar of CO₂. The grey dashed lines highlight features of the DFT spectra for comparison with the experimental spectra. The experimental spectra were taken 79° from the IL surface.

The IL + CO₂ simulated N K edge NEXAFS has a split π^* feature with peaks at 399.9 and 402.3 eV (compared to 399.9 and 401.7 eV in the experimental spectra). The split π^* peak arises from transitions from N^{2,r} only, the N atom that remains unreacted upon CO₂ absorption

(Figure S15). The σ^* peak occurs at a higher energy in the simulated spectrum (408.7 eV) compared to the experimental spectrum (407.7 eV). The σ^* peak has a slightly greater contribution from $N^{1,r}$, the N atom bonded to carbamate, compared to $N^{2,r}$.

We can compare the theoretical N K edge NEXAFS in Figures S14 and S15 to the unoccupied antibonding orbitals of [benzim]⁻ shown in Figures 4 (b) and (c) in the main text. From the $1\pi^*$ orbital we can infer that there are an equal number of excitations from N^1 and N^2 into the $1\pi^*$ orbital formed around atoms N^1 , C^4 , C^5 and N^2 . However, the introduction of CO_2 results in an asymmetry of this orbital compared to the same orbital in [benzim]⁻ pre-exposure. We can assign the two π^* peaks at 399.9 and 401.7 eV in the IL as presented experimental spectra as $N^{1,2} \rightarrow 1\pi^*$ and $N^{1,2} \rightarrow 2\pi^*$ transitions, respectively. Upon exposure to CO_2 , the peak at 399.9 eV is assigned to $N^{2,r} \rightarrow 1\pi^*$ transitions, while the peak at 401.7 eV is assigned to $N^{1,r}$ and $N^{2,r} \rightarrow 2\pi^*$ transitions.

Anion Orientation

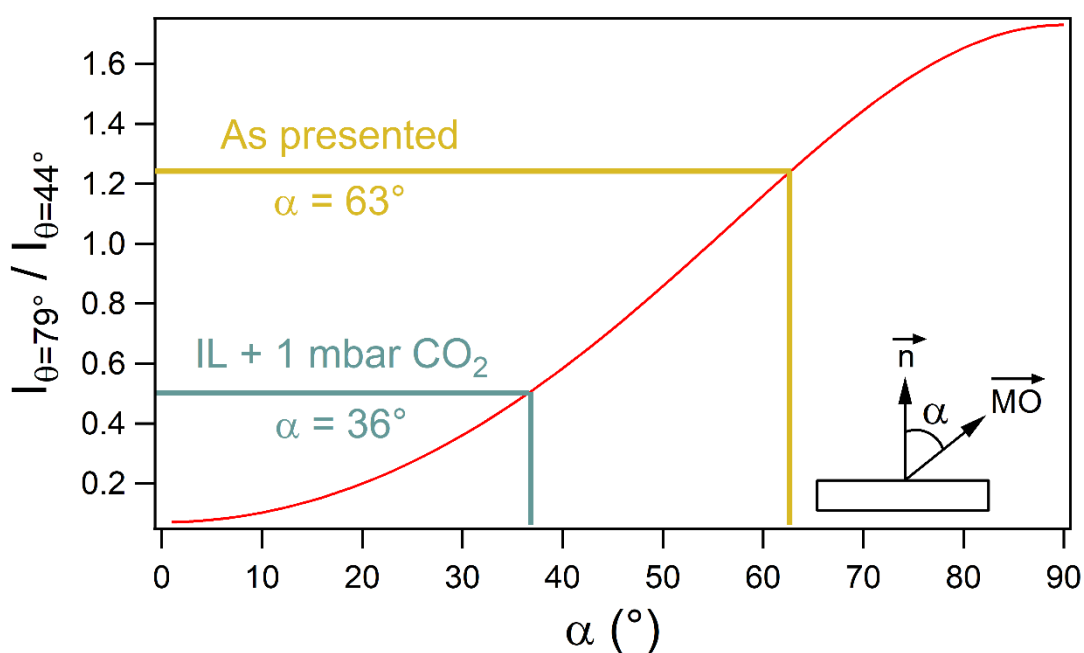


Figure S17: Intensity ratios measured at X-ray angles of incidence ($\vartheta=79^\circ$ and $\vartheta=44^\circ$) calculated as a function of the angle between the surface normal and the molecular orbital, α , using the Stöhr equations. Comparing these to the experimental intensity ratios gives $\alpha = 63^\circ$ for the as presented stage and $\alpha = 36^\circ$ upon exposure to 1 mbar of CO_2 .

The angle between the surface normal and the direction of the π^* molecular orbital is α . The ring of the anion is perpendicular to the π^* molecular orbital, therefore, the tilt of the benzim ring is given by $90^\circ - \alpha$. The angle α can be calculated using the Stöhr equations and the intensity ratio of the π^* resonance peaks at 285 eV in Figure 3.⁸ In the as presented case, the intensities of the π^* peaks are 0.97 and 0.78 for $\vartheta = 79^\circ$ and $\vartheta = 44^\circ$, respectively. This results in a resonance intensity ratio ($I_{\vartheta=79^\circ}/I_{\vartheta=44^\circ}$) of 1.24. When exposed to 1 mbar of CO_2 , these intensities are 0.59 and 1.19, respectively, resulting in an intensity ratio of 0.50. Comparing

these to the theoretical intensity ratio calculated as a function of α using the Stöhr equations (Figure S17), we can estimate the tilt of the benzim ring to be 27° from the surface normal in the as presented case. Upon exposure to CO₂, the benzim ring reorientates to 54° from the surface normal. We assume a random azimuthal orientation of the anion and a value of P = 0.85 for the polarisation factor.

References

- (1) Syres, K. L.; Thomas, A. G.; Flavell, W. R.; Spencer, B. F.; Bondino, F.; Malvestuto, M.; Preobrajenski, A.; Grätzel, M. Adsorbate-Induced Modification of Surface Electronic Structure: Pyrocatechol Adsorption on the Anatase TiO₂ (101) and Rutile TiO₂ (110) Surfaces. *J. Phys. Chem. C* **2012**, *116*, 23515–23525.
- (2) Taneichi, D.; Haneda, R.; Aramaki, K. A Novel Modification of an Alkanethiol Self-Assembled Monolayer with Alkylisocyanates to Prepare Protective Films against Copper Corrosion. *Corros. Sci.* **2001**, *43*, 1589–1600.
- (3) Ferrara, A. M.; Boufi, S.; Battaglini, N.; Botelho do Rego, A. M.; Reivilar, M. Hybrid Systems of Silver Nanoparticles Generated on Cellulose Surfaces. *Langmuir* **2010**, *26*, 1996–2001.
- (4) Stevens, J. S.; De Luca, A. C.; Pelendritis, M.; Terenghi, G.; Downes, S.; Schroeder, S. L. M. Quantitative Analysis of Complex Amino Acids and RGD Peptides by X-Ray Photoelectron Spectroscopy (XPS). *Surf. Interface Anal.* **2013**, *45*, 1238–1246.
- (5) Henderson, Z.; Thomas, A. G.; Wagstaffe, M.; Taylor, S. F. R.; Hardacre, C.; Syres, K. L. Reversible Reaction of CO₂ with Superbasic Ionic Liquid [P₆₆₆₁₄][Benzim] Studied with in Situ Photoelectron Spectroscopy. *J. Phys. Chem. C* **2019**, *123*, 7134–7141.
- (6) Wagstaffe, M.; Jackman, M. J.; Syres, K. L.; Generalov, A.; Thomas, A. G. Ionic Liquid Ordering at an Oxide Surface. *ChemPhysChem* **2016**, *17*, 3430–3434.
- (7) Cremer, T.; Stark, M.; Deyko, A.; Steinrück, H. P.; Maier, F. Liquid/Solid Interface of Ultrathin Ionic Liquid Films: [C₁C₁Im][Tf₂N] and [C₈C₁Im][Tf₂N] on Au(111). *Langmuir* **2011**, *27*, 3662–3671.
- (8) Stöhr, J. *NEXAFS Spectroscopy*; Springer: Berlin, 2003.
- (9) Walsh, J. F.; Dhariwal, H. S.; Gutiérrez-Sosa, A.; Finetti, P.; Muryn, C. A.; Brookes, N. B.; Oldman, R. J.; Thornton, G. Probing Molecular Orientation in Corrosion Inhibition via a NEXAFS Study of Benzotriazole and Related Molecules on Cu(100). *Surf. Sci.* **1998**, *415*, 423–432.
- (10) Grillo, F.; Garrido Torres, J. A.; Treanor, M. J.; Larrea, C. R.; Götze, J. P.; Lacovig, P.; Früchtl, H. A.; Schaub, R.; Richardson, N. V. Two-Dimensional Self-Assembly of Benzotriazole on an Inert Substrate. *Nanoscale* **2016**, *8*, 9167–9177.

- (11) Ehlert, C.; Holzweber, M.; Lippitz, A.; Unger, W. E. S.; Saalfrank, P. A Detailed Assignment of NEXAFS Resonances of Imidazolium Based Ionic Liquids. *Phys. Chem. Chem. Phys.* **2016**, *18*, 8654–8661.
- (12) Mercy, M.; Rebecca Taylor, S. F.; Jacquemin, J.; Hardacre, C.; Bell, R. G.; De Leeuw, N. H. The Addition of CO₂ to Four Superbase Ionic Liquids: A DFT Study. *Phys. Chem. Chem. Phys.* **2015**, *17*, 28674–28682.
- (13) Thomason, M. J.; Seabourne, C. R.; Sattelle, B. M.; Hembury, G. A.; Stevens, J. S.; Scott, A. J.; Aziz, E. F.; Schroeder, S. L. M. Self-Association of Organic Solutes in Solution: A NEXAFS Study of Aqueous Imidazole. *Faraday Discuss.* **2015**, *179*, 269–289.