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Title	Synthesis of Mesoporous Silica@Co-Al Layered Double Hydroxide Spheres: Layer-by-Layer Method and Their Effects on the Flame Retardancy of Epoxy Resins
Type	Article
URL	https://clock.uclan.ac.uk/11649/
DOI	https://doi.org/10.1021/am503412y
Date	2014
Citation	Jiang, Shu-Dong, Bai, Zhi-Man, Tang, Gang, Song, Lei, Stec, Anna A, Hull, T Richard, Hu, Yuan and Hu, Wei-Zhao (2014) Synthesis of Mesoporous Silica@Co-Al Layered Double Hydroxide Spheres: Layer-by-Layer Method and Their Effects on the Flame Retardancy of Epoxy Resins. ACS Applied Materials & Interfaces, 6 (16). pp. 14076-14086. ISSN 1944-8244
Creators	Jiang, Shu-Dong, Bai, Zhi-Man, Tang, Gang, Song, Lei, Stec, Anna A, Hull, T Richard, Hu, Yuan and Hu, Wei-Zhao

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<https://doi.org/10.1021/am503412y>

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Publication Information: Jiang, S.-D., Bai, Z.-M., Tang, G., Song, L., Stec, A.A., Hull, T.R., Hu, Y., Hu, W.-Z. *Synthesis of mesoporous silica@Co-Al layered double hydroxide spheres: Layer-by-layer method and their effects on the flame retardancy of epoxy resins*, ACS Applied Materials and Interfaces, 6 (16), pp. 14076-14086, (2014).

Synthesis of mesoporous silica@Co-Al layered double hydroxide spheres: layer-by-layer method and their effects on the flame retardancy of epoxy resins

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Abstract

Hierarchical mesoporous silica@Co-Al layered double hydroxide (m-SiO₂@Co-Al LDH) spheres were prepared through a layer by layer assembly process, in order to integrate their excellent physical and chemical functionalities. TEM results depicted that, due to the electrostatic potential difference between m-SiO₂ and Co-Al LDH, the synthetic m-SiO₂@Co-Al LDH hybrids exhibited that m-SiO₂ spheres were packaged by the Co-Al LDH nanosheets. Subsequently, the m-SiO₂@Co-Al LDH spheres were incorporated into an epoxy resin (EP) matrix to prepare specimens for investigation of reduced fire hazard behavior. Cone results indicated that m-SiO₂@Co-Al LDH incorporated obviously improved fire retardant of EP. A plausible flame-retardant mechanism was speculated on the basis of the analyses of thermal conductivity, char residues and pyrolysis fragments. Labyrinth effect of m-SiO₂ and graphitized carbon char formation catalyzed by Co-Al LDH play key roles in the flame retardance enhancement.

Keywords: mesoporous silica@Co-Al layered double hydroxide spheres, layer-by-layer method, polymer composites, thermal stability, flame retardancy, mechanism

1. Introduction

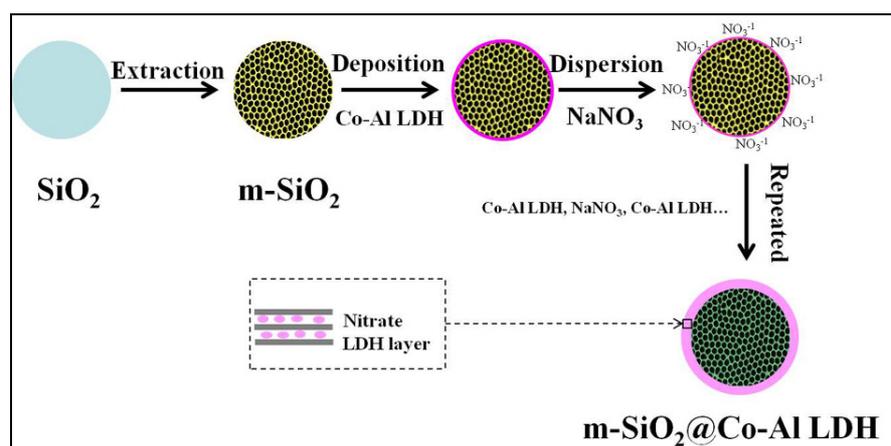
Mesoporous silica (m-SiO₂) has attracted much interest for its widespread applications in molecular adsorption, catalysis, gas and biological sensors, due to its high surface area, tunable pore size and very narrow pore size distribution.¹⁻³ The large-area and open-pore surface of m-SiO₂ is suitable for anchoring particles, trapping molecules, fast transport and optical manipulation.³⁻⁶ To date, m-SiO₂ has been further used to fabricate polymer nanocomposites with more excellent performance.⁷⁻⁹ Compared to other reported inorganic fillers,¹⁰⁻¹³ m-SiO₂ shows obvious advantages. On one hand, the mesopore sizes are much larger than the width of EP (about 0.5 nm). Therefore, EP can be easily penetrated into the mesochannels due to capillary force. Also, EP wrapped by the surface of m-SiO₂ exhibits high thermal stability. On the other hand, unlike nanosheets and nanotubes, the one-dimensional mesopore space is randomly oriented inside the composites to endow isotropy for nanocomposites.¹⁴ Although m-SiO₂ is clearly a promising polymer additive, few studies have reported the new application of m-SiO₂ on the flame retardancy of polymers.¹⁵ Thus, broadening the application of m-SiO₂ is a promising avenue of research.

As a kind of important lamellar materials, layered double hydroxides are currently obtained intense research interest owing to their excellent physical and chemical properties, which result in their extensive applications in catalysis, adsorption and flame retardant.¹⁶⁻¹⁹ However, the use of layered double hydroxides alone remains unsatisfactory to meet specific requirements for applications. Recently, core-shell structured hybrids have received tremendous attention in recent research for the combined characteristics of cores and shells. Shao et al. provided a novel Fe₃O₄@SiO₂@NiAl-LDH for the practical purification of recombinant proteins.²⁰ Silica microspheres decorated with CdS nanocrystals were synthesized by a facile process and used for the physical adsorption and direct photolysis of N-containing dyes.²¹ Teng et al. reported the fabrication of Fe₃O₄@SiO₂ with superparamagnetism, high magnetization and large surface area.²² Inspired by these, this study aims to combine of m-SiO₂ and Co-Al layered double hydroxide (Co-Al LDH) to improve their flame retardancy effectiveness. Co-Al LDH is a well-investigated member of the LDH family due to its ease of synthesis and delamination.²³ The exfoliated single sheets of Co-Al layered double hydroxide (Co-Al LDH) is positively charged with a thickness of 0.8 nm.²⁴ It is known that the surface of m-SiO₂ is negatively charged above the isoelectric point, which favors a layer coating of the positively charged material.²⁵ Electrostatic interactions between negatively charged m-SiO₂ and positively charged LDH sheets can create layered structures on the surface of m-SiO₂ through a layer-by-layer assembly method.

EP is one of the most widely exploited reactive polymeric resins in laminating, adhesive, coating and casting fields.²⁶⁻²⁷ However, its high flammability greatly restricted its application in some areas. Herein, m-SiO₂@Co-Al LDH spheres were synthesized by ultrasound assisted direct layered assembly of Co-Al LDH nanosheets on the surface of m-SiO₂ spheres, as schematically depicted in Scheme 1. Subsequently, the m-SiO₂@Co-Al LDH spheres were incorporated into

an EP matrix to prepare specimens for investigation of reduced fire hazard behavior. The synergistic effect between m-SiO₂ and Co-Al LDH on the thermal stability and fire resistance of the as-prepared nanocomposite was systematically investigated using thermogravimetric analysis (TGA) and cone calorimetry. Meanwhile, the analyses of thermal conductivity, char residue and pyrolysis fragments were utilized in an attempt to gain insight into flame-retardant mechanism.

Scheme 1. The synthetic route of m-SiO₂@Co-Al LDH.



2. Experimental

Tetraethyl orthosilicate (TEOS), cetyl trimethyl ammonium bromide (CTAB), ammonium hydroxide (28%), ethanol, Co(NO₃)₂ · 6H₂O, Al(NO₃)₃ · 9H₂O, NaNO₃, tetrahydrofuran (THF) 4,4'-diaminodiphenylmethane (DDM) were purchased from Sinopharm Chemical Reagent Co., Ltd. (Shanghai, China). Bisphenol-A type epoxy resin was commercially supplied by Hefei Jiangfeng Chemical Industry Co., Ltd., China.

2.1. Preparation of m-SiO₂

m-SiO₂ spheres were synthesized through a modified Stöber method. Typically, CTAB was dissolved in ethanol aqueous solution containing ammonium hydroxide. Then, TEOS (1 mL) was quickly injected into the above solution at 45 °C under strong mechanical stirring at a speed of 1500 rpm. The reactant molar ratio was 1.00 TEOS:0.0954 CTAB:2.84 NH₃:723 H₂O:125 C₂H₅OH. After 24 h of thermostatic reaction, the white precipitate was isolated by centrifugation at 5000 rpm for 10 min and cleaned by three cycles of centrifugation/washing/redispersion in ethanol. The yield of the final product was about 70%.

2.2. Preparation of Co-Al LDH

Co-Al LDH was prepared by a slow co-precipitation followed by ultrasound treatment. Briefly, 0.8 mL of aqueous ammonia solution (12.3 mmol) was added into 20 mL of deionized water to form solution A. Then 19.2 mL of an aqueous solution containing 0.96 mmol of Co(NO₃)₂ · 6H₂O

and 0.48 mmol of $\text{Al}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ was dropped at a constant rate of 40 mL/h into solution A under vigorous mechanical stirring with a speed of 2000 rpm (JJ-1 motor stirrer, 60W, 3000 rpm, Changzhou saipu Instruments Equipment Co., Ltd., China) followed by ultrasonication for 1 h. For the ultrasonication, this was done with a typical immersion bath. The model sonic bath was KQ-250 with a frequency of 40 KHz (Gongyi yuhua Instruments Equipment Co., Ltd., China). The obtained solid was collected and washed with pure water and ethanol. The yield of the Co-Al LDH was about 80%

2.3. Preparation of m-SiO₂@Co-Al LDH through the layer-by-layer method

m-SiO₂ (0.5 g) was dispersed in a Co-Al LDH formamide solution (0.1 g/100 mL). Then, the mixture was ultrasonically agitated for 20 min to facilitate the adsorption of Co-Al LDH nanosheets onto the m-SiO₂ surface. The particles were collected by centrifugation at 6000 rpm for 10 min and washed with deionized water. In the next step, the particles were dispersed in a NaNO₃ solution (100 mL, 2 g/L). The pink precipitate was isolated by centrifugation (6000 rpm, 10 min). m-SiO₂ coated with 10 layer pairs of nitrate and LDH nanosheets ($\text{NO}_3^{-1}/\text{LDH}$)₁₀ were prepared through repeating the above procedures 10 times. The yield of the final product was about 90%.

2.4. Preparation of EP/m-SiO₂@Co-Al LDH composite

A typical preparation of epoxy composite containing 2 wt% m-SiO₂@Co-Al LDH is illustrated below: m-SiO₂@Co-Al LDH (0.2 g) was added in THF with ultrasonication (Immersion sonic bath, KQ-250, 40 KHz, Gongyi yuhua Instruments Equipment Co., Ltd., China) to form a pink suspension. Then, DDM (1.6 g) was melt at 95 °C and mixed with epoxy resin (8.2 g) by hand with a spatula for 10 min. Subsequently, the mixture of epoxy resin and DDM was poured into the above suspension under magnetic stirring with a speed of 1000 rpm for 30 min until a homogeneous mixture was formed. The mixture was placed in a vacuum chamber at 60 °C for 12 h to remove THF. Finally, the EP/m-SiO₂@Co-Al LDH composite was cured at 100 °C for 2 h and post cured at 150 °C for 2 h. After curing, the composite was permitted to cool to room temperature. For the preparation of neat EP, EP/Co-Al LDH and EP/m-SiO₂ composites, a analogous prepared procedure was adopted except the variation of the additives.

2.5. Characterization

Powder X-ray diffraction (XRD) measurements were performed on a Japan Rigaku DMax/rA rotating anode X-ray diffractometer, using Cu K α radiation ($\lambda = 0.154$ nm) at 40 kV. Transmission electron microscopy (TEM), Energy dispersive X-ray spectroscopy (EDX) and high-resolution transmission electron microscopy (HRTEM) analyses were performed using a Hitachi model H-800 TEM with an accelerating voltage of 200 kV. Before measurement, samples were dispersed in ethanol followed by ultrasonication for 30 min at room temperature. The homogeneous mixtures were dripped on carbon-coated copper grids. UV-vis absorption spectra were detected by a Solid3700 (SHIMADZU) spectrometer. X-ray photoelectron spectroscopy (XPS) spectra were recorded on an ESCALAB 250 X-ray photoelectron spectrometer employing a monochromatic Al K α X-ray source. Thermogravimetric analyses (TGA) of samples were implementing with the aid of a Q5000 thermal analyzer (TA Co., USA) at a heating speed of 20 °C min⁻¹ in nitrogen atmosphere. The combustion properties of EP and EP

composites were carried out on a cone calorimeter based on ASTM E1354/ISO 5660. Every specimen with the sizes of $100 \times 100 \times 3 \text{ mm}^3$ wrapped in an aluminum foil was exposed horizontally to a heat flux of 35 kW/m^2 . Three parallel runs were performed for each sample to obtain averages. Microstructures of the char residues were studied by a JEOL JSM-2010 field-emission scanning electron microscopy (FESEM). Raman spectroscopy measurements were carried out with a SPEX-1403 laser Raman spectrometer (SPEX Co, USA) with excitation provided in back-scattering geometry by a 514.5 nm argon laser line. Direct pyrolysis-mass spectrometry (DP-MS) analysis was performed with a Micromass GCT-MS spectrometer using the standard direct insertion probe for solid polymer materials, at a heating rate of $15 \text{ }^\circ\text{C min}^{-1}$ in the range of 30-700 $^\circ\text{C}$.

3. Result and Discussion

3.1. Characterization of m-SiO₂@Co-Al LDH

TEM image in Figure 1A shows that m-SiO₂ spheres are monodispersed with a uniform diameter of around 500 nm. High-magnification TEM image depicts that the mesochannels of the spheres are continuous throughout the shell with openings at surface and are radially oriented to the sphere surface (Figure 1B), which can absorb EP molecular chain. The surface of m-SiO₂ spheres is negatively charged as indicated by a zeta potential of -45 mV . Hence, deposition of the oppositely charged Co-Al LDH precursors via electrostatic interaction can easily occur without adding any additives or binding agents. The morphology of m-SiO₂@Co-Al LDH was further examined by TEM (Figures 2C, D and E). It can be observed that the final composite consists of well dispersed spheres with an average diameter of around 550 nm. The surface of m-SiO₂ spheres (dark colour) is coated with a layer of lighter coloured LDH nanosheets. As shown in Figure 1F, the building block of nanosheets shows a repeating fringe of $\sim 0.8 \text{ nm}$. They are assembled with a face-to-face stacking to form a layer structure with a lateral scale of 8 nm. The zeta potential of m-SiO₂@Co-Al LDH was measured to be $+26.7 \text{ mV}$, further supporting that positively charged LDH nanosheets have been effectively coated on the surface of m-SiO₂. The existence of Co-Al LDH coating on the surface of m-SiO₂ was also identified by EDX, as shown in Figure 1G.

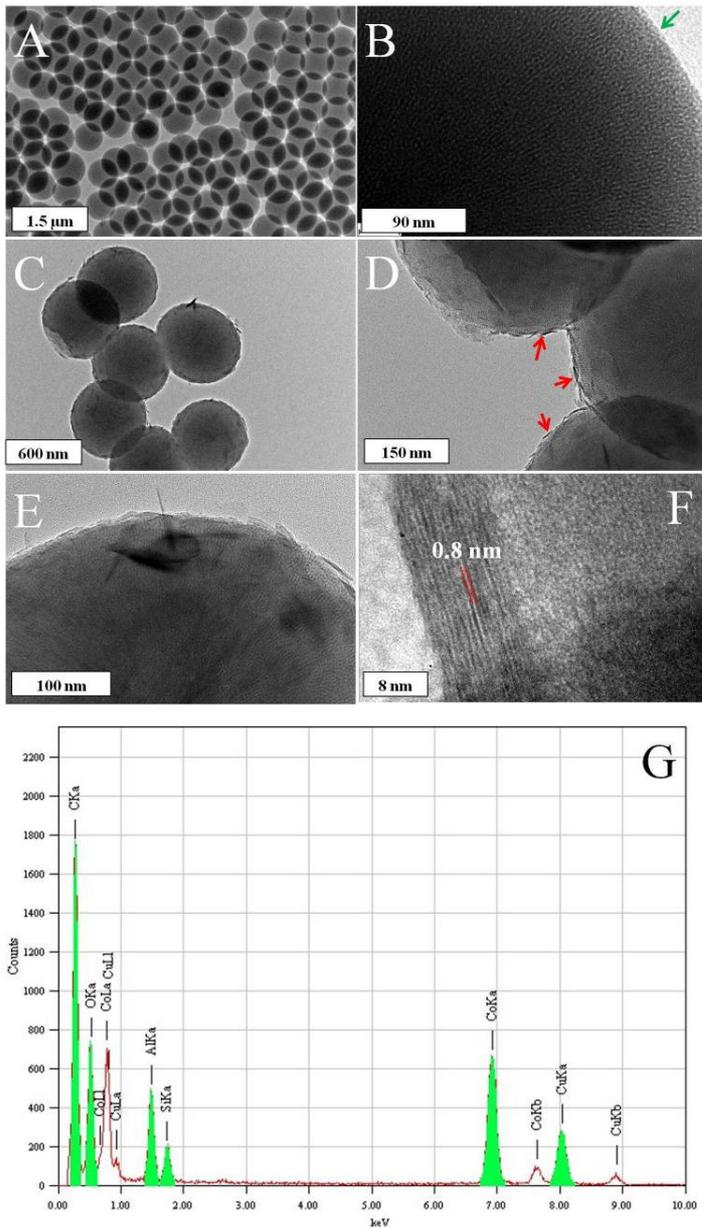


Figure 1. TEM images of m-SiO₂ (A, B), m-SiO₂@Co-Al LDH (C, D, E), HRTEM image (F) of m-SiO₂@Co-Al LDH and EDX spectra (G) of m-SiO₂@Co-Al LDH.

As far as the layer number is concerned, the morphology of samples appears nearly same. For the blank m-SiO₂, the particles exhibit microspheres with a smooth surface (Figure 2A). With the layer number increased to 10, the spherical morphology of monodisperse m-SiO₂ spheres is preserved well after the deposition of nitrate LDH shell (Figure 2B). The only noticeable difference between the spheres with or without shells is the surface roughness. When the layer number is increased to 20, the surface of SiO₂ spheres (dark colour) is coated with more LDH

nanosheets and become more coarser (Figure 2C). Moreover, there is an enhanced absorbance at $\lambda=525$ nm (Figure 2D) with the increase of the layer number, which is in accordance with the color change of the samples.

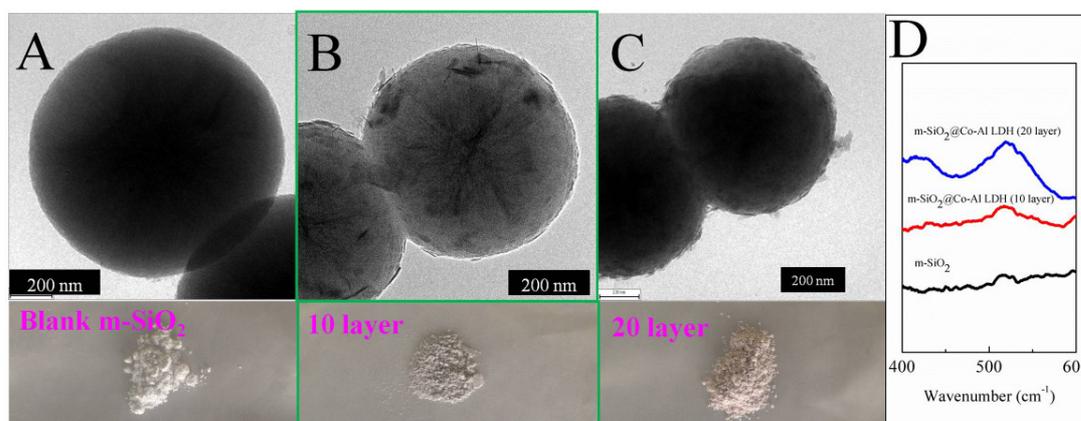


Figure 2. TEM images, digital photos (A, B, C) and UV-vis absorption spectra (D) of m-SiO₂, m-SiO₂@Co-Al LDH (10 layers) and m-SiO₂@Co-Al LDH (20 layers).

The composition and structure of m-SiO₂, Co-Al LDH and m-SiO₂@Co-Al LDH were identified by XRD, as shown in Figure 3. The wide-angle XRD pattern of m-SiO₂ shows a broad peak around 21°, revealing its amorphous nature.²⁸ For Co-Al LDH, all diffraction peaks can be indexed as Co-Al LDH.²⁹⁻³⁰ The broadened diffraction peaks of Co-Al LDH indicate that the synthetic sample is composed of nanocrystals.³¹ Compared the m-SiO₂, Co-Al LDH and m-SiO₂@Co-Al LDH wide-angle XRD patterns, two relatively weak peaks of m-SiO₂@Co-Al LDH at around can be assigned to (003) and (012) of layered Co-Al LDH structure (Figure 3A). Moreover, the (006) peak of Co-Al LDH and the broad peak between 20° and 25° of m-SiO₂ were overlapping, so the broad peak of m-SiO₂ may cover the weak (006) peak of Co-Al LDH in the m-SiO₂@Co-Al LDH structure. The small-angle XRD diffraction of m-SiO₂ and m-SiO₂@Co-Al LDH exhibit a typical diffraction pattern of hexagonally packed mesopores, associated with the 100, 110, and 200 reflections of hexagonal symmetry with the space group p6mm (Figure 3B). This is in good agreement with that of the as-made bulk mesoporous silica MCM-41, suggesting a similar ordered hexagonal mesostructure.³² The d_{100} spacing of the mesostructure is 3.7 nm, corresponding to a unit cell parameter of 4.3 nm. Nitrogen sorption isotherm of the m-SiO₂@Co-Al LDH spheres depicts a type IV curve according to the IUPAC nomenclature with a sharp capillary condensation step and a hysteresis loop in the p/p_0 range 0.2–0.4 (Figure 3C), revealing characteristics of mesoporous materials with narrow pore size distribution. The surface area and pore volume are calculated to be as high as about 1020 m² g⁻¹ and 0.80 cm³ g⁻¹. The detailed pore size distribution calculated based on the nonlocal density functional theory reveals the spheres have uniform micropores of about 1.2 nm (Figure 3D).

XPS analysis was used to investigate the surface composition and chemical state of the m-SiO₂@Co-Al LDH spheres, which could provide information of the interaction between m-SiO₂ and Co-Al LDH. In Figure 4A it can be seen that the product contains Si, O, Co and Al elements, indicating the formation of m-SiO₂@Co-Al LDH. The Co_{2p} spectrum is shown in Figure 4B. Two peaks corresponding to Co2p_{3/2} and Co2p_{1/2} are centered at 781.2 eV and 796.8 eV, respectively. The Co2p_{3/2} peak of m-SiO₂@Co-Al LDH appears at higher position than that of normal Co-Al

LDH, indicating a stronger interaction between silica shell and Co-Al LDH species.³³⁻³⁵ The presence of m-SiO₂ can also be clearly confirmed from the Si2p curve (Figure 4C). The position of 102.06 eV compares well with literature reports on binding energies for SiO₂. Furthermore, a dark-field scanning transmission electron microscopy (STEM) analysis of m-SiO₂@Co-Al LDH was performed to obtain more detailed information about the structure. As shown in Figure 4D, the dark-field STEM clearly demonstrates that the as-synthesized m-SiO₂@Co-Al LDH have a typical core-shell structure, and EDX element mapping of the same particles further shows the spatial distributions of Si, O, Co and Al in nanoparticles of m-SiO₂@Co-Al LDH. The strong Si and O signal across the sphere confirms the m-SiO₂ core, while the Co and Al signals both detected in the surface region clearly suggests the adsorption of Co-Al LDH particles.

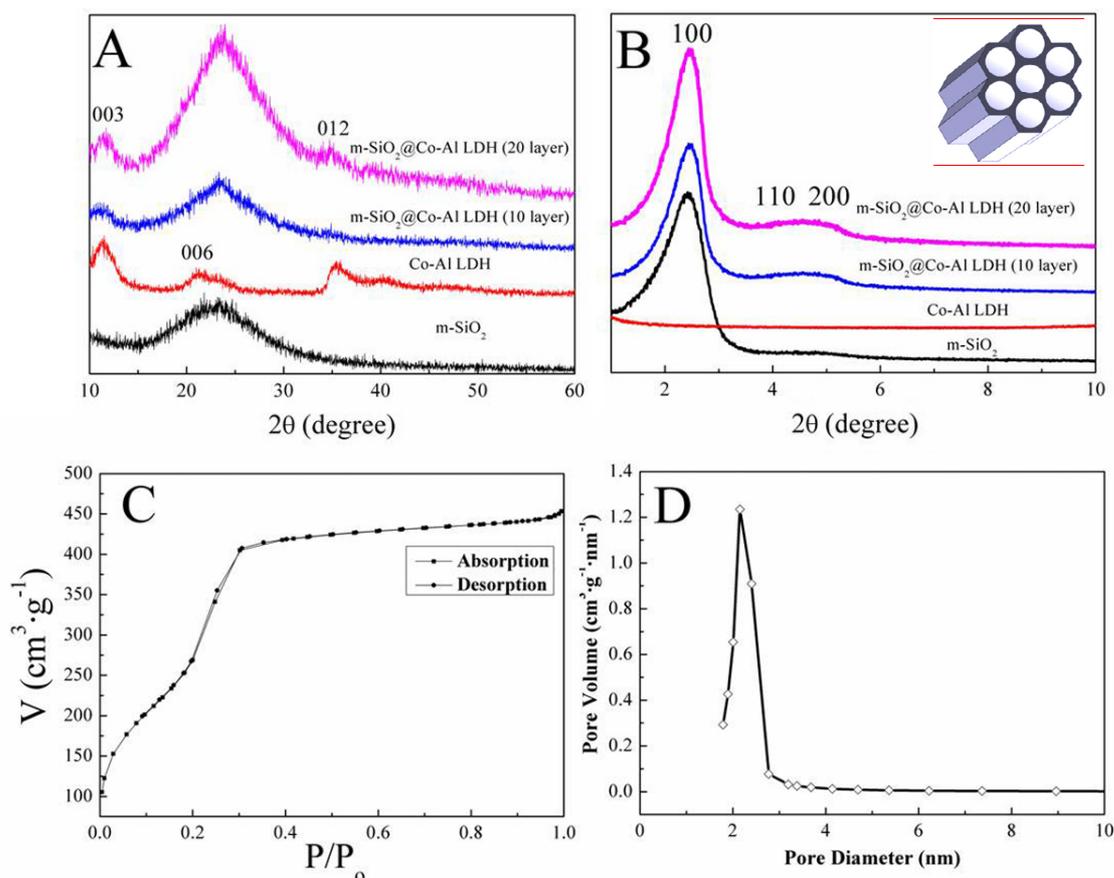


Figure 3. The wide-angle XRD pattern (A) and the low-angle XRD pattern (B) m-SiO₂@Co-Al LDH, nitrogen sorption isotherm (C) and the pore size distribution curve of m-SiO₂@Co-Al LDH (D). Inset in (B) is a structural model of the mesoporous silica shells, showing hexagonal symmetry.

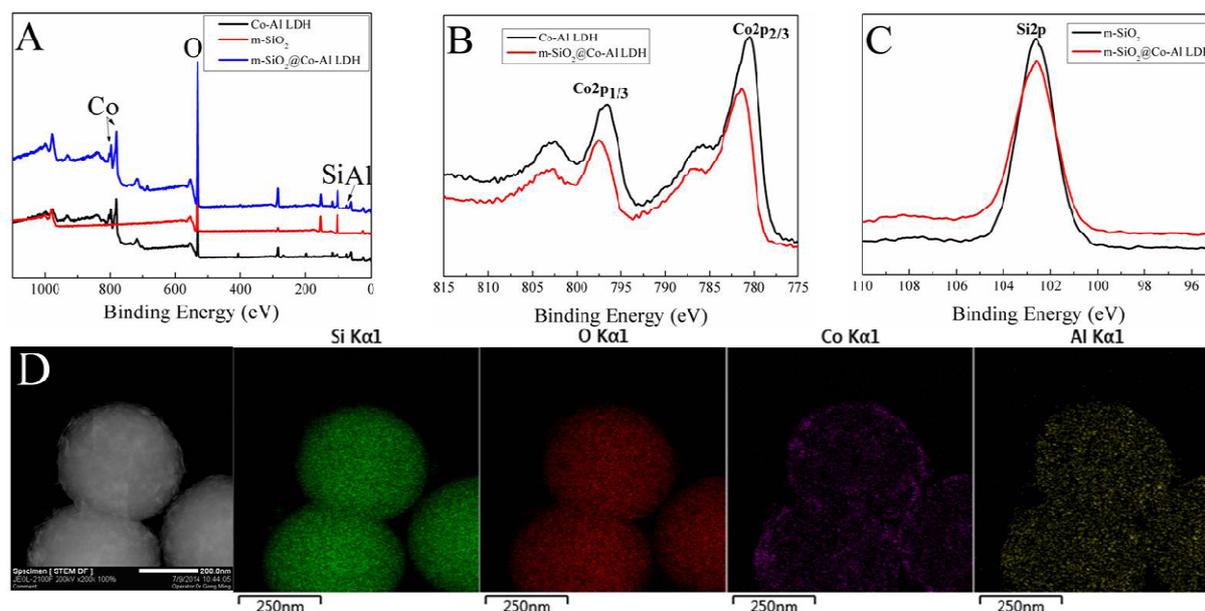


Figure 4. XPS survey spectra of Co-Al LDH, m-SiO₂ and m-SiO₂@Co-Al LDH (A), Co 2p XPS spectra of Co-Al LDH and m-SiO₂@Co-Al LDH (B), Si 2p XPS spectra of m-SiO₂ and m-SiO₂@Co-Al LDH (C), dark-field STEM image and elemental mapping of m-SiO₂@Co-Al LDH (D).

3.2. Thermal degradation of EP and its nanocomposites

The thermal stability of m-SiO₂, Co-Al LDH, m-SiO₂@Co-Al LDH, EP and its nanocomposites was evaluated by TGA in nitrogen atmosphere. Figure 5 shows the TGA and derivative thermogravimetric (DTG) curves, and the corresponding data are given in Table 1. The temperatures at which 10% ($T_{-10\%}$), 50% ($T_{-50\%}$) and maximum (T_{max}) mass loss occurs are used as the measure of initial degradation temperature, half degradation temperature and maximum degradation temperature, respectively. Pure m-SiO₂ shows a three-step thermogravimetric profile corresponding to the loss of physically adsorbed water (room temperature–120 °C), the polycondensation of the silica network (120–210 °C) and the further condensation and dehydration of silanol groups (210–350 °C).^{36–38} Co-Al LDH exhibits a three-step thermogravimetric profile accompanied a mass loss of about 34.3 %, which should be due to surface-adsorbed water (from room temperature to 200 °C), chemisorbed water (200–350 °C) and the water arising from the dehydroxylation of the layers (225–500 °C).¹⁹ In comparison, m-SiO₂@Co-Al LDH displays a consecutive mass loss with a weight loss of about 14.3 %, implying that the integration of Co-Al LDH with m-SiO₂ obviously enhances the thermal stability of Co-Al LDH. From Figure 5 and Table 1, EP/m-SiO₂ nanocomposite is less thermally stable than pure EP when evaluated by $T_{-10\%}$, which is possibly attribute to the catalysis degradation initiated by silanol groups attached on m-SiO₂ as brønsted acid sites.³⁹ At higher temperatures, the EP molecular chains encapsulated by the silica walls shows high thermal stability due to the barrier effect of its porous structure, leading to a delayed degradation process of EP.⁴⁰ It is already known that transition metals including Co can cause catalytic degradation of polymer.⁴¹ As expected, the Co-Al LDH incorporated results in an earlier degradation process of EP. Also, the

introduction of m-SiO₂@Co-Al LDH leads to the decrease of degradation temperature, which is attributable to the earlier degradation of EP triggered by the catalytic activity of Co-Al LDH. However, compared to EP/Co-Al LDH, EP/m-SiO₂@Co-Al LDH nanocomposite presents the improved thermal stability owing to the barrier effect of m-SiO₂. As far as the char yield is concerned, the incorporation of Co-Al LDH or m-SiO₂@Co-Al LDH leads to the increase of the char residues at 700 °C, due to the catalytic carbonization effect of LDH.⁴² Additionally, from the DTG profiles, the maximum mass loss rates (the peak of DTG curve) of EP/m-SiO₂ and EP/m-SiO₂@Co-Al LDH composites are much lower than that of pure EP, indicating that m-SiO₂ plays an effective mass-transport barrier role.

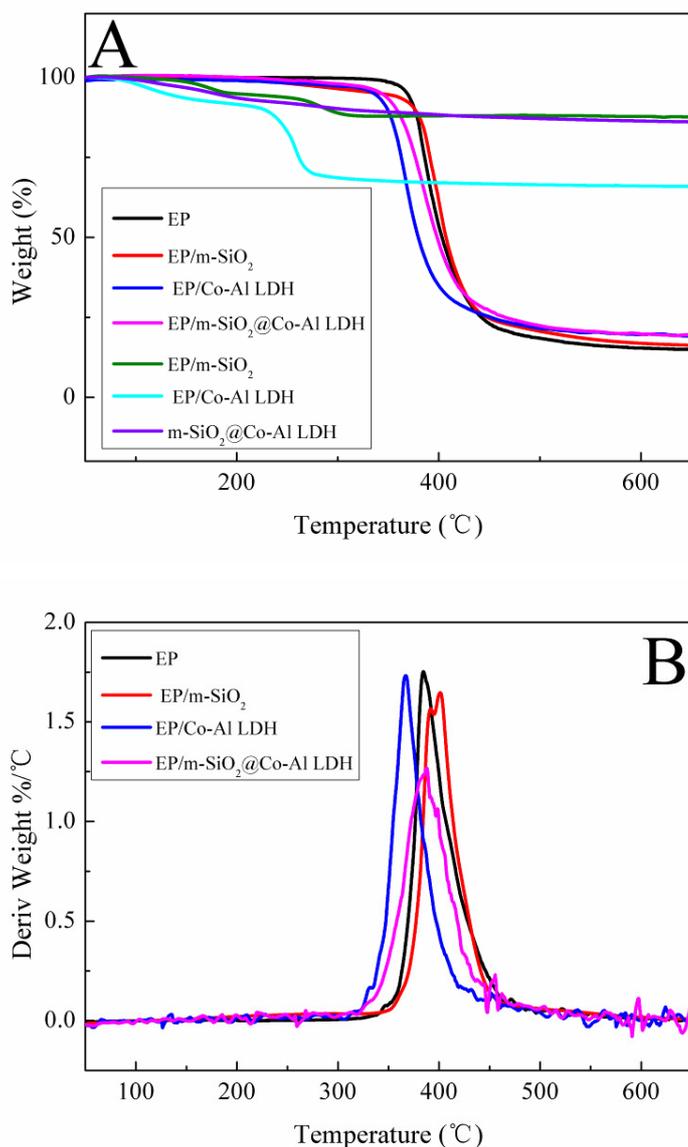


Figure 5. TGA (A) and DTA (B) curves of pure EP and its nanocomposites in nitrogen.

Table 1. TGA data for EP and its nanocomposites in nitrogen.

Sample	$T_{-10\%}$ (°C)	$T_{-50\%}$ (°C)	T_{max} (°C)	Char (%)
EP	376.0	403.2	385.2	14.7
EP/m-SiO ₂	376.4	407.1	401.8	15.8
EP/Co-Al LDH	349.0	379.6	366.5	18.4
EP/m-SiO ₂ @Co-Al LDH	358.1	398.3	389.1	19.0

3.3. Fire properties of EP and its composites

Cone calorimetry is a widely used tool for investigating flammability of materials to forecast fire behaviour in real-world fire.⁴³ Heat release rate (HRR) curves for EP and its composites are shown in Figure 6 and some important parameters obtained from cone calorimetry are listed in Table 2. With the incorporation of inorganic additives, the peak heat release rate (pHRR), total heat release (THR), effective heat of combustion (EHC), total smoke release (TSR) and maximum average heat rate emission (MAHRE) values are shifted to lower values except that the THR value of EP/m-SiO₂ increases. This phenomenon is attributed to the prolonged burning time as can be seen from the broad peak width in the pHRR curve and the formation of the combustible fragment from the catalytic degradation effect of m-SiO₂.⁴⁴⁻⁴⁷ Compared to pure EP, the addition of m-SiO₂@Co-Al LDH brings about a 39.3% maximum decrease in pHRR, a 36.2% maximum decrease in THR, a 15.8% maximum decrease in EHC, a 23.8% maximum decrease in TSR and a 16.7% maximum decrease in MAHRE. The results exhibit that the EP/m-SiO₂@Co-Al LDH nanocomposite presents the best flame retardancy among all the nanocomposites. According to the earlier literatures, m-SiO₂ has an interconnected porous structure with low thermal conductivity.⁴⁸ Heat and mass transport pathways in such a pore structure are long and tortuous, resulting in the diffusion of heat and the emission of volatile degradation products restricted. Thus, the labyrinth effects from m-SiO₂, containing mass barrier and thermal barrier effects, are hypothesized to be the main reasons for the flame retardancy enhancement. The thermal barrier effect limits the permeation of heat; meanwhile the mass barrier effect inhibits the escape of degradation products. In addition, the improved char yield of EP/m-SiO₂@Co-Al LDH on the basis of TGA results is another possible factor for the enhanced flame retardancy.

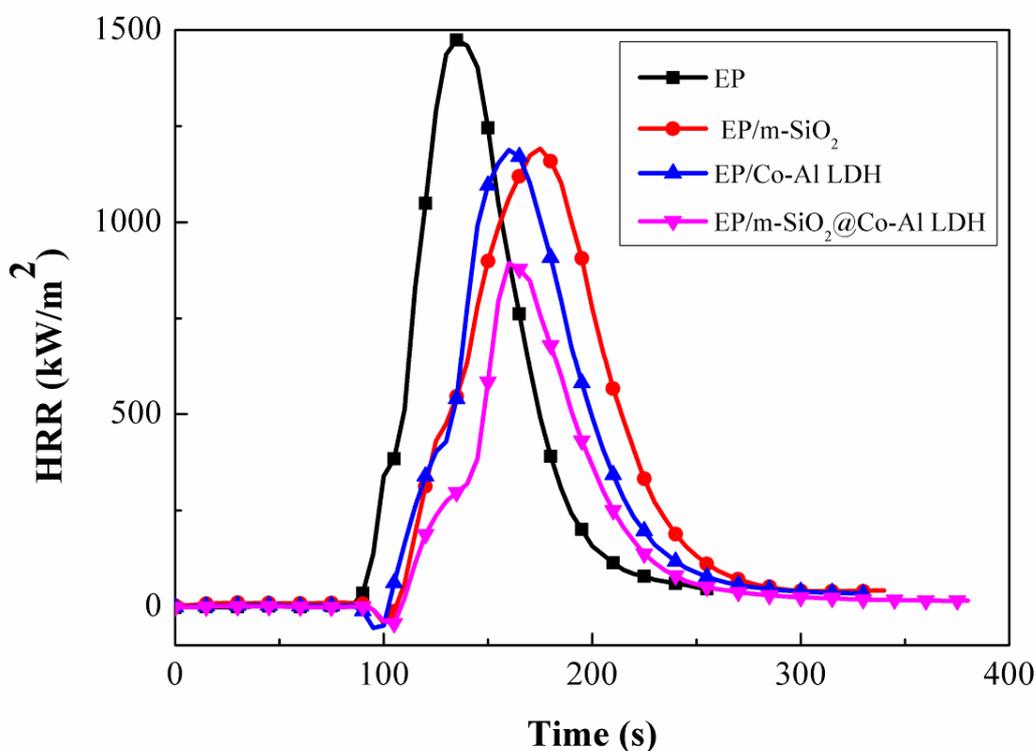


Figure 6. HRR curves of EP and its nanocomposites.

Table 2. Cone calorimetry data of EP and its nanocomposites.

Sample	<i>pHRR</i> (kW/m ²)	<i>THR</i> (MJ/m ²)	<i>EHC</i> (MJ/Kg)	<i>MARHE</i> (kW/m ²)	<i>TSR</i> (m ² /m ²)	Char yield %
EP	1473.4	87.8	24.1	434.2	2734.9	5.2 %
EP/m-SiO ₂	1191.2	96.5	26.4	383.8	2291.8	7.1 %
EP/Co-Al LDH	1188.1	84.3	23.8	421.6	2370.3	12.1 %
EP/m-SiO ₂ @Co-Al LDH	893.8	56.0	20.3	361.7	2085.2	20.7 %

3.4. Flame retardant mechanism

To understand the flame-retardant mechanism, we studied the thermal conductivity, the char residues and the degradation products of EP and its nanocomposites. Figure 7 shows the thermal conductivity of the nanocomposites at different temperatures. The thermal conductivity of the EP/Co-Al LDH is around 0.4131 W/m·K and it changes slightly with the increase of the temperature. As for the EP/m-SiO₂@Co-Al LDH nanocomposite, the thermal conductivity decreases compared to the EP/Co-Al LDH (0.4131 W/m·K) to 0.2789 W/m·K with a

filler concentration of 2 wt%. It is suggested that heat diffusion was more difficult in the presence of m-SiO₂@Co-Al LDH. Combined with the BET results (Figures 3C and D), m-SiO₂@Co-Al LDH with larger surface area makes the path length longer that the transfer of oxygen and combustible products was inhibited, leading to retarded heat and mass release. Therefore, it is reasonable to believe that the low thermal conductivity and the labyrinth effect contributed by m-SiO₂@Co-Al LDH could lead to the better flame resistance.

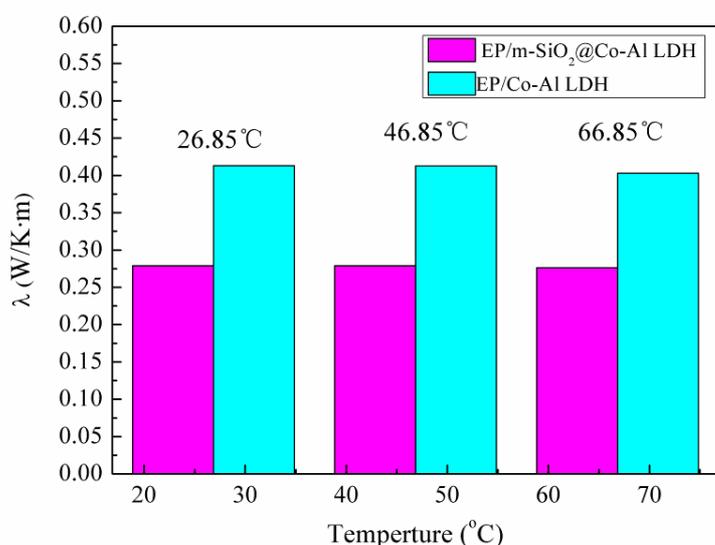


Figure 7. Thermal conductivity plot of the nanocomposites at different temperatures.

Figure 8A shows images of the char residues after cone calorimetry test, in which there are little residues left after thermal degradation of pure EP. The incorporation of m-SiO₂@Co-Al LDH results in the significant improvement of the char yield, corresponding to the TGA results. To further explore the flame-retardant mechanism, Raman spectroscopy was utilized to characterize the structure and component of the residues. The Raman spectra of the residues of EP and EP/m-SiO₂@Co-Al LDH (Figure 8B) depicts two bands at 1356 cm⁻¹ (D-band) and 1591 cm⁻¹ (G-band), which are associated with the vibration of the carbon atoms in disordered graphite or glassy carbons and the carbon atoms in crystalline graphite., respectively.⁴⁹ This result indicates the formation of the graphitized char during EP combustion, due to the catalysis of m-SiO₂@Co-Al LDH.⁵⁰ As well known, the graphitized char is more conducive to suppress the diffusion of heat and mass during pyrolysis than disordered graphite or glassy carbons, which is ascribed to its higher thermal stability and compactness. Moreover, Figures 8C, D show the typical microstructure of the residues of EP and EP/m-SiO₂@Co-Al LDH nanocomposite. It is clearly found that a continuous and cohesive char surface is formed after the EP/m-SiO₂@Co-Al LDH nanocomposite combustion. The residue with a more cohesive and compact layer is beneficial for the inhibition of the heat, mass and oxygen exchange, thereby improving the flame retardancy.

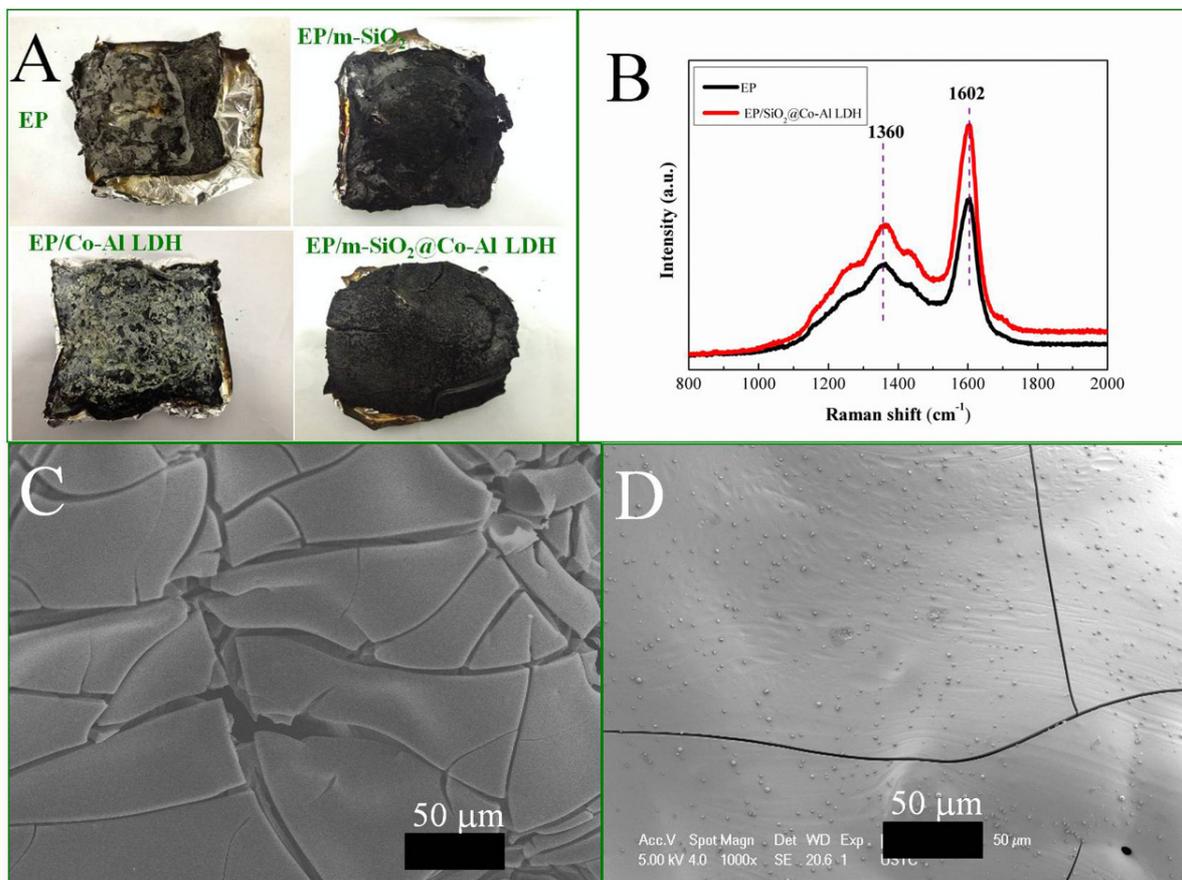


Figure 8. Digital photos (A) and Raman spectra (B) of the residues from EP and its nanocomposites, SEM images of residues of EP (C) and EP/m-SiO₂@Co-Al LDH (D).

It is well-known that the pyrolysis products act as a key role in the flame-retardant performance of polymers. For the purpose of understanding the flame-retardant mechanism of the EP nanocomposites, the DP-MS tool was used to analyze the pyrolysis products. Figure 9 depicts the total ion current (TIC) chromatogram of EP and EP/m-SiO₂@Co-Al LDH nanocomposite, and EI-MS spectra corresponding to the TIC peaks with the maximum intensity. The pyrolysis fragment ions with the molecular weight of x are labelled as M_x and identified in Figure 10. For pure EP, the strong peak at 255 m/z corresponds to $C_{17}H_{19}O_2$. Also, the peaks at 269 m/z , 213 m/z , 197 m/z and 135 m/z can be assigned to the pyrolysis products of $C_{17}H_{21}N_2O$, $C_{14}H_{13}O_2$, $C_{15}H_{18}$, and $C_9H_{11}O$. Additionally, some fragment ions can regroup to produce new products in high temperature. Such as, the benzene and aniline reconstitute to form carbazole (166 m/z), and the recombination between carbazole and benzene reconstitute generates a few polycyclic aromatic hydrocarbons (497 m/z). In contrast, the main fractions in the degradation products of the EP/m-SiO₂@Co-Al LDH include $C_{14}H_{13}O_2$ (213 m/z), which is with lower carbon numbers than that of neat EP. Actually, previous literatures reported that solid acids can catalyze the degradation of polymer, which results in the formation of pyrolysis products with lower carbon numbers, and the degradation products with lower carbon numbers could be easily catalyzed carbonization in the presence of metal oxides.⁵¹⁻⁵² m-SiO₂ has been reported to be a most

efficient solid acid for catalytic degradation of polymer due to the presence of many acid sites.⁵³⁻⁵⁴ In this work, carbon numbers of pyrolysis products from the EP/m-SiO₂@Co-Al LDH composite is also decreased compared with that of EP. Thus, this undoubtedly results from the catalytic degradation effect for EP by m-SiO₂.⁵³⁻⁵⁴ Also, degradation products with lower carbon numbers could extend contacting time of with metal oxides catalyst under the labyrinth effect of m-SiO₂. With different Co catalysts, Gong et al., used a simple method to synthesize carbon nanospheres through the carbonization of polystyrene.⁵⁵ Therefore, it is reasonable to believe that the enhanced flame resistant property for EP/m-SiO₂@Co-Al LDH is attributed to the synergism of m-SiO₂ and Co-Al LDH. In view of the results of the volatile pyrolysis fragment ions, the mechanism for the improved fire resistant property of EP/m-SiO₂@Co-Al LDH nanocomposite is explained as follow. During the combustion process, m-SiO₂ with catalytic activity leads to the formation of pyrolysis products with lower carbon numbers, which can be easily catalyzed carbonization in the presence of metal oxides. Meanwhile, Co-Al LDH can catalyze carbonization of degradation products. Moreover, m-SiO₂ plays as a barrier which can absorb degraded products to extend contacting time of with metal compound catalyst. Furthermore, the degraded products are dehydrogenated and catalytically converted into char by the combination of m-SiO₂ labyrinth effect and Co-Al LDH catalysis effect.

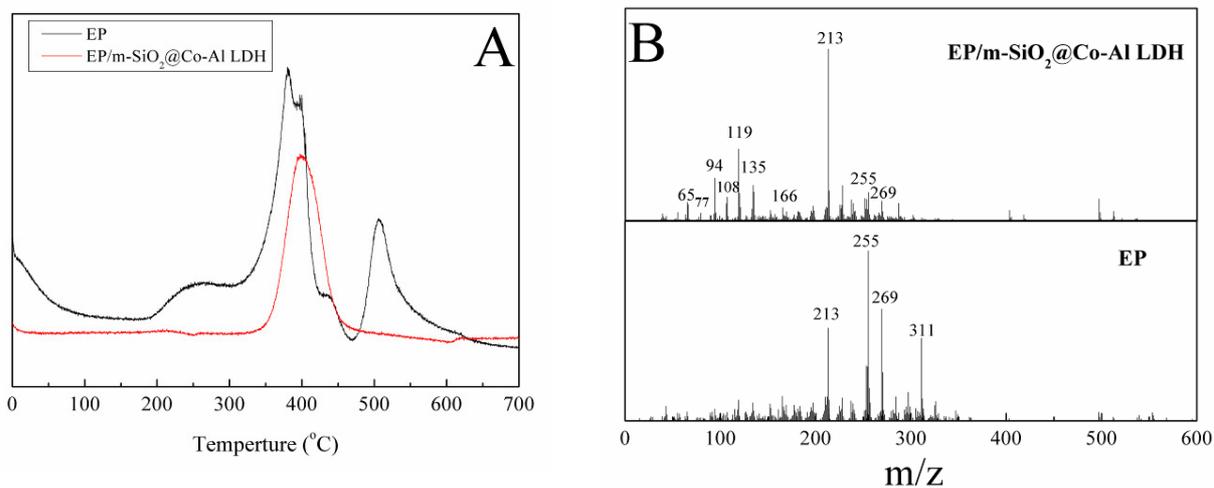


Figure 9. Total ion current (TIC) curves (A) of the decomposition process of EP and EP/m-SiO₂@Co-Al LDH, and EI-MS spectra (B) of compounds evolved from EP and EP/m-SiO₂@Co-Al LDH at the maximum of the peak in the TIC curves.

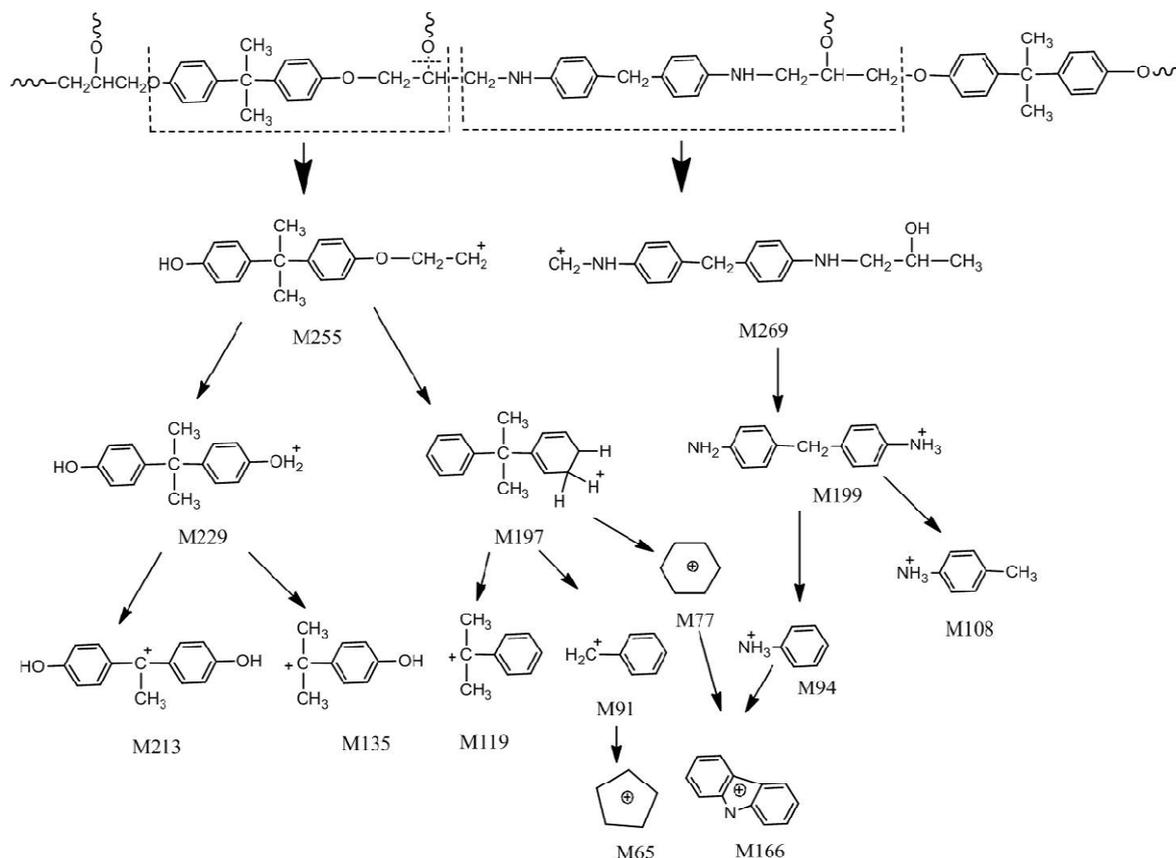


Figure 10. Simplified mass fragmentations of the EP/m-SiO₂@Co-Al LDH nanocomposite.

4. Conclusions

In conclusion, m-SiO₂@Co-Al LDH spheres were synthesized by ultrasound assisted direct layered assembly of Co-Al LDH nanosheets on the surface of m-SiO₂ spheres, and its composition and structure was identified by XRD, BET and XPS. The morphological characterization showed that, owing to the electric potential difference between m-SiO₂ and Co-Al LDH, m-SiO₂@Co-Al LDH exhibited that m-SiO₂ spheres were packaged by the Co-Al LDH nanosheets. Incorporation of 2 wt% m-SiO₂@Co-Al LDH into EP led to the increase of the char yield and the decrease of DTG peak value. Moreover, the pHRR, THR, EHC, TSR and MAHRE values for EP/m-SiO₂@Co-Al LDH were obviously reduced. A plausible flame-retardant mechanism was speculated based on the analyses of thermal conductivity, char residue and pyrolysis fragments. m-SiO₂ plays as a mass and heat barrier to prevent the diffusion of heat and the spread of the pyrolysis products, and hence prolong the contact time between the pyrolysis products and the Co-Al LDH catalyst. The significant reduction of the fire hazard was primarily due to the synergistic action between the labyrinth effect of m-SiO₂ and the catalytic effect of Co-Al LDH.

Acknowledgements

The work was financially supported by the National Natural Science Foundation of China (21374111) and the National Basic Research Program of China (973 Program) (2012CB719701).

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